

Review – Thermodynamics Processes

□ According to a law of $PV^n = \text{constant}$

Process	Description	Result of Ideal Gas Law
<i>isochoric</i>	constant volume ($V_1 = V_2$)	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
<i>isobaric</i>	constant pressure ($P_1 = P_2$)	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
<i>isothermal</i>	constant temperature ($T_1 = T_2$)	$P_1V_1 = P_2V_2$
<i>polytropic</i>	-none-	$\frac{P_1}{P_2} = \left(\frac{V_2}{V_1}\right)^n = \left(\frac{T_1}{T_2}\right)^{\frac{n}{n-1}}$
<i>isentropic</i>	constant entropy ($S_1 = S_2$)	

A **polytropic** process is a thermodynamic process that obeys the relation $pV^n=C$: where n is the polytropic index. If $n=\gamma$, the process is **isentropic**.

In thermodynamics, an **adiabatic** process is a type of thermodynamic process that occurs without transferring heat or mass between the thermodynamic system and its environment. Unlike an isothermal process, an adiabatic process transfers energy to the surroundings only as work.

Thermodynamic Processes Test Review

Didier Musso



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