

THERMODYNAMICS

Thermal equilibrium and definition of temperature (zeroth law of thermodynamics), heat, work and internal energy. First law of thermodynamics, isothermal and adiabatic processes. Second law of thermodynamics: reversible and irreversible processes

Thermodynamics: - Thermodynamics is a branch of physics that deals with the concept of heat and temperature, and the transmission of heat and other forms of energy.

System: Anything bounded by a closed surface is called a system.

Surrounding: - Anything outside the system which has a direct effect on the system is called surrounding.

There are 3 classes of systems:

1. **Open system:** It is a system which can exchange both matter and energy with the surrounding.
2. **Closed system:** A system which can exchange only energy, not matter with the surrounding is called closed system.
- 3: **Isolated system:** is a system which is not influenced by its surroundings.

Thermodynamic variables or Co-ordinates or State variables

Thermal state of a thermodynamic system is determined by its *temperature, pressure and volume*. These quantities are called thermodynamic variables. The process by which the system goes from one thermodynamic state to another state is called *thermodynamic process*.

Thermodynamic variables are of 2 types:

1. **Extensive variable:** It indicates the size of the system. Eg. Volume, mass, internal energy.
2. **Intensive variables:** They do not indicate the size of the system. Eg: temperature, pressure, density.

Work: - In thermodynamics we consider only the external work which involves the *interaction between the system and its surroundings*. Any internal work done by one part of a system on another part is *not considered* in thermodynamics. In thermodynamics work done *by a system is taken as positive and work done on the system is taken as negative*. In thermodynamics work is associated with change in volume.

Work done by a thermodynamic system

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