

●32. A 20.0 g lump of chromium has a heat capacity of 9.03 J/°C. What is the specific heat of chromium?

$$c_{\text{sp Cr}} = 0.4515 \text{ J/g } ^\circ\text{C} \quad (= 0.452 \text{ J/g } ^\circ\text{C})$$

●33. What is the heat capacity of a 100. g bar of silver?

$$C_{\text{Ag}} = 23.5 \text{ J/}^\circ\text{C}$$

●34. If a spool of copper wire has a heat capacity of 13.6 J/°C, what is the spool's mass?

$$m_{\text{Cu}} = 35.32 \text{ g} \quad (= 35.3 \text{ g})$$

●35. A 30.0 g gold coin and a 30.0 g silver coin sitting on a window sill each receive 63.0 J of energy from the Sun. If they were originally at the same temperature, which will be warmer after the energy is absorbed?

$$\otimes t_{\text{Au}} = +16.3 \text{ }^\circ\text{C}; \quad \otimes t_{\text{Ag}} = +8.94 \text{ }^\circ\text{C}; \quad \text{The gold coin will be warmer.}$$

●36. A 425 g glass jar has a temperature of 20.0 °C. If $c_{\text{sp glass}} = 0.837 \text{ J/g } ^\circ\text{C}$, what is the jar's temperature after it absorbs 3560 J?

$$= 30.0 \text{ }^\circ\text{C}$$

●37. A 50.0 g asbestos pad's temperature rises from 25.0 °C to 27.5 °C when $1.05 \cdot 10^5 \text{ J}$ of heat is absorbed. [Data revised. See Errata file.]

a. What is the pad's heat capacity?

$$42000 \text{ J/}^\circ\text{C}$$

b. What is the specific heat of asbestos?

$$c_{\text{sp}} = 840 \text{ J/g}^\circ\text{C}$$

●38. A 10.0 g lump of lead is placed in boiling water (100.0 °C) until its temperature is constant. Then the lead is removed, quickly dried, and placed into 10.0 g of water at 20.0 °C in an adiabatic calorimeter. The final temperature of the lead and water is 22.4 °C.

a. How much heat does the water gain?

$$Q_{\text{water}} = +100 \text{ J} \quad (= +1.0 \cdot 10^2 \text{ J})$$

b. How much heat does the lead lose?

$$Q_{\text{Pb}} = -100 \text{ J} \quad (= -1.0 \cdot 10^2 \text{ J})$$

c. What is the lump's heat capacity?

$$1.28 \text{ J/}^\circ\text{C}$$

d. What is the specific heat of lead?

$$.128 \text{ J/g } ^\circ\text{C}$$

●39. 25.0 g of boiling water is added to a calorimeter near room temperature (20.0 °C). The final temperature of the calorimeter and water is 81.5 °C.

a. How much heat does the water lose?

$$Q_{\text{w}} = -1933 \text{ J} \quad (= -1930 \text{ J})$$

b. What is the calorimeter's heat capacity?

$$31.43 \text{ J/}^\circ\text{C}$$

●40. How much heat does 1.00 g of steam/water release as it cools from 105 °C to 55.0 °C?

$$Q_{\text{total}} = -2454 \text{ J} \quad (= -2450 \text{ J})$$

●41. A 10.0 g lump of metal at 77.6 °C is dropped into a calorimeter containing 10.0 g of ice and 10.0 g of water at 0.0 °C. 1.0 gram of the ice melts before thermal equilibrium is reestablished.

a. What is the final temperature of the mixture?

b. How much heat does the ice that melts gain?

$$Q_{\text{ice}} = +333 \text{ J} \quad (= +330 \text{ J})$$

c. How much heat do the water and calorimeter gain?

d. What is the lump's heat capacity?

$$C_{\text{metal}} = 4.29 \text{ J/}^\circ\text{C} \quad (= 4.3 \text{ J/}^\circ\text{C})$$

e. What is the metal's specific heat?

$$c_{\text{sp metal}} = 0.429 \text{ J/g } ^\circ\text{C} \quad (= 0.43 \text{ J/g } ^\circ\text{C})$$

●42. In order to double the rate of radiant energy output of a hot plate, how much hotter must the plate be? Express your answer to three significant digits.

$$T_2 = 1.197$$

●●43. 50.0 g of liquid oxygen at its boiling point is sealed in a 50.0 g adiabatic aluminum

Thermodynamics Mixed Review Answers

Paul Painter



Thermodynamics Mixed Review Answers:

Quantum and Statistical Thermodynamics with Spectroscopic Applications Mr. Rohit Manglik,2024-03-04

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proposes for the first time a systematic framework to describe athermal fluctuation developing stochastic thermodynamics for non Gaussian processes while thermal fluctuations are mainly addressed from the viewpoint of Gaussian stochastic processes in most of the conventional studies First the book provides an elementary introduction to the stochastic processes and stochastic thermodynamics The author derives a Langevin like equation with non Gaussian noise as a minimal stochastic model for athermal systems and its analytical solution by developing systematic expansions is shown as the main result Furthermore the author shows a thermodynamic framework for such non Gaussian fluctuations and studies some thermodynamics phenomena i e heat conduction and energy pumping which shows distinct characteristics from conventional thermodynamics The theory introduced in the book would be a systematic foundation to describe dynamics of athermal fluctuation quantitatively and to analyze their thermodynamic properties on the basis of stochastic methods

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