

Question number	Answer	Marks	Guidance
6 (c)	$M(\text{NaNO}_3) = 85.0 \text{ g mol}^{-1}$ $n(\text{NaNO}_3) = 0.04 \text{ mol}$ $n(\text{O}_2) = 0.02 \text{ mol}$ Volume of $\text{O}_2 = 0.0200 \times 24\,000 = 480 \text{ cm}^3$	B1 B1 B1 B1	
7 (a)	$0.0250 \times 23.0 = 0.575 \text{ g}$	B1	
7 (b)	$2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$	B1	
7 (c)	$n(\text{H}_2) = 0.0125 \text{ mol}$ Volume $\text{H}_2 = 0.0125 \times 24\,000 = 300 \text{ cm}^3$	B1 B1	
7 (d) (i)	$n(\text{NaOH}) = n(\text{Na}) = 0.0250 = c \times 50/1000$ $c = 0.500 \text{ mol dm}^{-3}$	B1	
7 (d) (ii)	$c = 0.500 \times 40.0 = 20.0 \text{ g dm}^{-3}$	B1	
8 (a)	58.5 g	B1	
8 (b)	$n(\text{Cl}_2) = (2.5 \times 10^3)/24 = 1.04 \times 10^2 \text{ mol}$ $n(\text{NaOH}) = 2 \times 1.04 \times 10^2 = 2.08 \times 10^2 \text{ mol}$ $2.08 \times 10^2 = 4.00 \times V \text{ (in dm}^3\text{)}$ $\therefore V = 5.2 \times 10^1 \text{ dm}^3$	B1 B1 B1	
9 (a) (i)	C : H = $54.55/12.0 : 9.09/1.0 : 36.36/16.0 = 4.55 : 9.09 : 2.27$ Empirical formula = $\text{C}_2\text{H}_4\text{O}$	B1 B1	
9 (a) (ii)	$pV = nRT$ $n = \frac{(103 \times 10^3) \times (72.0 \times 10^{-6})}{8.314 \times 373} =$ 0.002 39 mol $M = \frac{0.2103}{0.00239} = 88.0$	B1 B1 B1 B1	
9 (a) (iii)	Molecular formula = $\text{C}_2\text{H}_4\text{O} \times 88/44 = \text{C}_4\text{H}_8\text{O}_2$	B1	
10 (a) (i)	$\text{Al}^{3+}$ $\text{SO}_4^{2-}$	B1 B1	

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**Michael Seilmaier**



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