

CHAPTER 1 Assessing and Managing Risk

STUDY GUIDE FOR CHAPTER 1 LESSON 1

The Highway Transportation System

- A.** The National Highway Safety Act controls the regulations for vehicle registration, road construction and maintenance, and driver licensing. What do you think the effects on the following would be if suddenly we had no National Highway Safety Act?

Roads and Highways: _____

Drivers and Pedestrians: _____

Cars and Trucks: _____

- B.** Although each state has its own laws governing the use of motor vehicles, the federal government has also set national highway and driving standards. Tell by marking S or F next to the provision whether it is a state or a federal law.

- _____ 1. Establishes 65 MPH as the maximum speed that may be driven anywhere.
- _____ 2. Sets national standards for all motor vehicle equipment.
- _____ 3. Regulates vehicle inspection.
- _____ 4. Enforces traffic laws.
- _____ 5. Tells what highway maintenance the state must provide.
- _____ 6. Regulates vehicle registration.
- _____ 7. Instructs automobile manufacturers about what safety devices they must provide.
- _____ 8. Assigns points to driver's record for traffic violations.
- _____ 9. Regulates driver licensing.
- _____ 10. Requires the vehicle manufacturer to recall a vehicle and correct any defects discovered after a vehicle is sold.

- C. FIND OUT MORE.** Interview somebody you know who is a driver. Ask that person what he or she does personally to reduce risk when driving. What does the driver do to keep the vehicle in good, safe condition? What does she or he do to anticipate the actions of others? Do the driver and all passengers wear seat belts? Has the driver ever driven while very tired or sick? What has the driver done in the last year to improve his or her driving skills? Write your findings on a separate piece of paper.

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction
 $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{C}_6\text{H}_4\text{COONa} + \text{H}_2\text{O}$
 the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

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