



### Honors Chemistry: Chapter 4- Problem Set (with some 6)

- All answers and work on a separate sheet of paper!

Classify the following as always true (AT), sometimes true (ST), or never true (NT)

1. Atoms of one element change into atoms of another element during chemical reactions.
2. Atoms combine in one-to-one ratios to form compounds.
3. Atoms of one element are different from atoms of other elements.
4. According to Dalton's atomic theory, atoms are composed of protons, electrons and neutrons.
5. Atoms of elements are electrically neutral.
6. The mass of an electron is equal to the mass of a neutron.
7. The charge on all protons is the same.
8. The atomic number of an element is the sum of the protons and electrons in an atom of that element.
9. An atom of nitrogen has 7 protons and 7 neutrons.
10. Relative atomic masses are expressed in amu's.
11. The number of neutrons in the nucleus can be calculated by subtracting the atomic number from the mass number.
12. In his periodic table, Mendeleev arranged the element in ascending order of atomic number.
13. The representative elements are the Group A elements.
14. The transition metals and inner transition metals are the Group B elements.
15. The element in Group 4A, period 3 is gallium.

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Match each description in Column B with the correct term in Column A

Column A	Column B
1. atom	a. the group 7A elements
2. Dalton's atomic theory	b. horizontal rows of the periodic table
	c. Group A elements
3. Democritus	d. arrangement of elements according to similarities in their properties
4. electrons	e. a portion of the B Group elements
	f. Group 1A elements
5. protons	g. Group 0 elements or group 8A
6. neutrons	h. vertical column on the periodic table
7. nucleus	i. represents the number of protons in an atom
8. atomic number	j. 1/12 of the mass of the carbon -12 atom
9. mass number	k. total number of protons and neutrons in an atom
10. isotopes	l. weighted average mass of the atoms in a naturally occurring sample of an element
11. atomic mass unit (amu)	m. atoms that have the same number of protons but different numbers of neutrons
12. atomic mass	n. positively charged subatomic particles
13. periodic table	o. subatomic particles with no charge
	p. negatively charged subatomic particles
14. periods	q. the central positively charged core of the atom
15. group	r. teacher in ancient Greece
16. representative elements	s. theory to explain the reactive behavior of elements
17. alkali metals	t. smallest particle of an element that retains the properties of that element
18. transition metals	
19. halogens	
20. noble gases	

Solve the following problems:

# Responsive Ed Chemistry Unit 4 Answer Key

**U.S. Nuclear Regulatory Commission**



## Responsive Ed Chemistry Unit 4 Answer Key:

Resources in Education ,2000-04      **Young Talents in Polymer Science** Alexander Böker, Frank Wiesbrock, 2018-07-05

This book is a printed edition of the Special Issue Young Talents in Polymer Science that was published in Polymers      EPA Publications Bibliography Quarterly Abstract Bulletin United States. Environmental Protection Agency, 1995-04      **EPA Publications Bibliography** United States. Environmental Protection Agency, 1995      Title List of Documents Made Publicly Available U.S. Nuclear Regulatory Commission, 1989      *El-Hi Textbooks & Serials in Print, 2000* ,2000      **AP® Chemistry Crash Course, 2nd Ed., Book + Online** Adrian Dingle, 2014-02-13

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