

Write the Perfect Resume in 13 Easy Steps

To write a perfect resume, you must read comprehensive guides from trusted resume writing resources. Don't worry, though—we've got you covered. Now, we present to you the anatomy of a perfect resume and the steps you should follow to write one.



Resume Writing Guide

J Elliott



Resume Writing Guide:

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have $\frac{21.3 \text{ g}}{32 \text{ g/mol}} = 0.666 \text{ mol}$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g/mol} = 63.9 \text{ g}$ c determine the percent yield $\frac{46.0 \text{ g}}{63.9 \text{ g}} \times 100 = 72.0\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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reacted with 8.6 g of H_2

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 $80 \text{ g } I_2O_5 \times \frac{1 \text{ mol } I_2O_5}{283.5 \text{ g } I_2O_5} \times \frac{1 \text{ mol } I_2}{2 \text{ mol } I_2O_5} \times \frac{253.8 \text{ g } I_2}{1 \text{ mol } I_2} = 28 \text{ g } I_2$

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 text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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