

Reversible Reactions and Chemical Equilibrium Worksheet

Supporting questions

- 1) Can you draw the reversible arrow symbol?
- 2) What do you observe when you heat up hydrated copper sulfate?
- 3) Define the term dynamic equilibrium?
- 4) Why can dynamic equilibrium only occur in a closed system?
- 5) What are the 3 factors that alter the position of equilibrium?
- 6) What happens to the position of equilibrium if you add a catalyst?
- 7) For the chemical reaction below can you say what happens to the concentration of ammonia if you (a) Add in more hydrogen (b) increase the pressure (c) increase the temperature (the forward reaction is **exothermic**).



Nitrogen Hydrogen Ammonia

- 8) Why might a high temperature still be used?

Thanks for taking the time to complete these questions, you can check your answers on the next page.

Reversible Reactions And Equilibrium Ch 18 2

Worksheet

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