

Chemistry 1020 Final Exam Review

1. A 25.0 mL sample of 1.50 M sodium chlorate (NaClO) is titrated with 1.00 M HCl . The K_a of HClO is 2.9×10^{-8} .

Weak Acid

(a) What is the pH after 20.0 mL of HCl has been added?



$$n_{\text{ClO}^-} = (1.50\text{M})(0.0250\text{L}) = 0.0375\text{mol}$$

$$n_{\text{H}_3\text{O}^+} = (1.00\text{M})(0.0200\text{L}) = 0.0200\text{mol}$$

buffer: b/c we are between start and equivalence point

$$\text{pH} = \text{pK}_a + \log \frac{[\text{ClO}^-]}{[\text{HClO}]} = \text{pK}_a + \log \frac{n_{\text{ClO}^-}}{n_{\text{HClO}}}$$

$$n_{\text{ClO}^-} \text{ remaining} = n_{\text{ClO}^-} - n_{\text{H}_3\text{O}^+} = 0.0375\text{mol} - 0.0200\text{mol} = 0.0175\text{mol}$$

$$n_{\text{HClO}} = n_{\text{H}_3\text{O}^+} = 0.0200\text{mol} \quad \text{pH} = -\log(2.9 \times 10^{-8}) + \log \left(\frac{0.0175\text{mol}}{0.0200\text{mol}} \right) = 7.49$$

(b) What is the pH at the equivalence point?

eq. point: all ClO^- consumed. n_{HClO} added = n_{ClO^-} from start



$$V_{\text{eq}} = n_{\text{eq}} \text{ added}$$

$$V_{\text{eq}} = \frac{n_{\text{eq}}}{C_{\text{eq}}}$$

$$= \frac{0.0375\text{mol}}{1.00\text{M}} = 0.0375\text{L}$$

$$[\text{HClO}] = \frac{n_{\text{HClO}}}{V_{\text{total}}} = \frac{0.0375\text{mol}}{0.0250\text{L} + 0.0375\text{L}}$$

$$= 0.600\text{M}$$

$$\text{pH} = -\log[\text{HClO}]$$

$$= -\log 0.600$$

(c) What is the pH after 40.0 mL of HCl have been added?

eq. V of HCl was 37.5 mL

$$V \text{ of HCl after eq. point} = 40.0\text{mL} - 37.5\text{mL} = 2.5\text{mL}$$

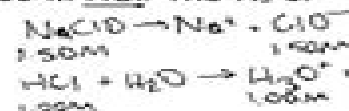
$$n_{\text{H}_3\text{O}^+} \text{ in solution} = (0.0025\text{L})(1.00\text{M}) = 0.0025\text{mol} = n_{\text{H}_3\text{O}^+}$$

$$[\text{H}_3\text{O}^+] = \frac{n_{\text{H}_3\text{O}^+}}{V_{\text{total}}} = \frac{0.0025\text{mol}}{0.0375\text{L}} = 0.0667\text{M}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+] = -\log(0.0667) = 1.18$$

Weak Acid

Strong Acid



C17

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**Zayika Wilson-Kennedy, Goldie S.
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Frierson**

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