

CHAPTER 15 REVIEW

Acid-Base Titration and pH**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Calculate the following values without using a calculator.

- _____ a. The $[\text{H}_3\text{O}^+]$ is 1×10^{-6} M in a solution. Calculate the $[\text{OH}^-]$.
- _____ b. The $[\text{H}_3\text{O}^+]$ is 1×10^{-9} M in a solution. Calculate the $[\text{OH}^-]$.
- _____ c. The $[\text{OH}^-]$ is 1×10^{-12} M in a solution. Calculate the $[\text{H}_3\text{O}^+]$.
- _____ d. The $[\text{OH}^-]$ in part c is reduced by half, to 0.5×10^{-12} M. Calculate the $[\text{H}_3\text{O}^+]$.
- _____ e. The $[\text{H}_3\text{O}^+]$ and $[\text{OH}^-]$ are _____ (directly, inversely, or not) proportional in any system involving water.

2. Calculate the following values without using a calculator.

- _____ a. The pH of a solution is 2.0. Calculate the pOH.
- _____ b. The pOH of a solution is 4.73. Calculate the pH.
- _____ c. The $[\text{H}_3\text{O}^+]$ in a solution is 1×10^{-3} M. Calculate the pH.
- _____ d. The pOH of a solution is 5.0. Calculate the $[\text{OH}^-]$.
- _____ e. The pH of a solution is 1.0. Calculate the $[\text{OH}^-]$.

3. Calculate the following values.

- _____ a. The $[\text{H}_3\text{O}^+]$ is 2.34×10^{-5} M in a solution. Calculate the pH.
- _____ b. The pOH of a solution is 3.5. Calculate the $[\text{OH}^-]$.
- _____ c. The $[\text{H}_3\text{O}^+]$ is 4.6×10^{-8} M in a solution. Calculate the $[\text{OH}^-]$.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.4. $[\text{H}_3\text{O}^+]$ in an aqueous solution = 2.3×10^{-3} M.

- _____ a. Calculate $[\text{OH}^-]$ in this solution.

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RM Cervero



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