

Resonance structures are multiple electronic representations of the same molecule. The only difference between resonance structures is the location of electrons.

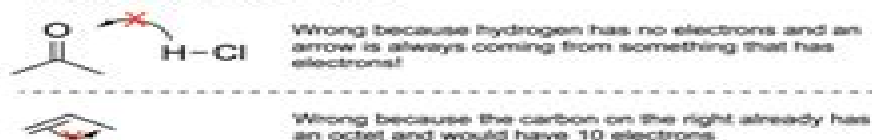
Curved Arrows represent the movement of an electron pair



Curved Arrows Rules:

1. Second row elements (C, N, O...) CANNOT exceed an octet (8 electrons)
2. An arrow should always start from a place that has electrons such as a lone pair or a bond

Wrong Curved Arrows



Transformation

Main Resonance Patterns

Double bond next to a positive charge



Draw one arrow from the double bond to the neighboring bond with positively charged atom.



Transformation

Double bond next to a lone pair



Draw one arrow from the lone pair to the neighboring bond and another arrow from the double bond to the atom.

Example:



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A lone pair next to a positive charge



Draw one arrow from the lone pair to the neighboring bond with the positively charged atom.

Note: The net charge of all resonance forms must be equal

Hint: Draw out all of the carbons, hydrogens and lone pairs of electrons when you begin to practice resonance

Relative importance of resonance forms

1. The more bonds in a resonance form, the more important it is.



2. The less charge, the better resonance is.



3. The more electronegative an atom is, the easier it is for it to bear a negative charge, and not positive charge.



oxygen is more electronegative than carbon and can bear the negative charge better.

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Resonance Structures Cheat Sheet

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Resonance Structures Cheat Sheet:

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