

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{3 \text{ O}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ Fe}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{312.07 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = 0.51 \text{ mol H}_3\text{PO}_4$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{1 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 10.89 \text{ g S}$$

$$\frac{12.75 \text{ g SO}_2}{64.06 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 19.13 \text{ g S}$$

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**Susan K. Johnsen, Joyce VanTassel-
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Current Challenges in Modeling Cellular Metabolism Daniel Machado,Kai H.

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Table of Contents Response Key To Ap Chemistry 2013 Exam

1. Understanding the eBook Response Key To Ap Chemistry 2013 Exam
 - The Rise of Digital Reading Response Key To Ap Chemistry 2013 Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Response Key To Ap Chemistry 2013 Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Response Key To Ap Chemistry 2013 Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Response Key To Ap Chemistry 2013 Exam
 - Personalized Recommendations
 - Response Key To Ap Chemistry 2013 Exam User Reviews and Ratings
 - Response Key To Ap Chemistry 2013 Exam and Bestseller Lists
5. Accessing Response Key To Ap Chemistry 2013 Exam Free and Paid eBooks
 - Response Key To Ap Chemistry 2013 Exam Public Domain eBooks
 - Response Key To Ap Chemistry 2013 Exam eBook Subscription Services
 - Response Key To Ap Chemistry 2013 Exam Budget-Friendly Options

6. Navigating Response Key To Ap Chemistry 2013 Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Response Key To Ap Chemistry 2013 Exam Compatibility with Devices
 - Response Key To Ap Chemistry 2013 Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Response Key To Ap Chemistry 2013 Exam
 - Highlighting and Note-Taking Response Key To Ap Chemistry 2013 Exam
 - Interactive Elements Response Key To Ap Chemistry 2013 Exam
8. Staying Engaged with Response Key To Ap Chemistry 2013 Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Response Key To Ap Chemistry 2013 Exam
9. Balancing eBooks and Physical Books Response Key To Ap Chemistry 2013 Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Response Key To Ap Chemistry 2013 Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Response Key To Ap Chemistry 2013 Exam
 - Setting Reading Goals Response Key To Ap Chemistry 2013 Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Response Key To Ap Chemistry 2013 Exam
 - Fact-Checking eBook Content of Response Key To Ap Chemistry 2013 Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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