

Stoichiometry Review

1. The combustion of propane occurs via the reaction:



How many grams of oxygen are required to burn completely 10.0 g of propane?

$$\left(\frac{10.0 \text{ g C}_3\text{H}_8}{44.09652 \text{ g/mol}} \right) \left(\frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right) \left(31.9988 \text{ g/mol} \right) = 36.3 \text{ g O}_2$$

2. Small quantities of chlorine can be prepared in the laboratory by the reaction:



What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

$$\left(\frac{100.0 \text{ g MnO}_2}{86.9368 \text{ g/mol}} \right) \left(\frac{1 \text{ mol Cl}_2}{1 \text{ mol MnO}_2} \right) \left(22.4 \text{ L/mol} \right) = 25.8 \text{ L Cl}_2$$

3. The most common ore of arsenic is mispickel, FeSAs. Upon heating this ore, free arsenic is obtained. How many grams of FeSAs are required to produce 10.0 g of arsenic according to the following reaction?



$$\left(\frac{10.0 \text{ g As}}{74.9216 \text{ g/mol}} \right) \left(\frac{1 \text{ mol FeSAs}}{1 \text{ mol As}} \right) \left(162.8346 \text{ g/mol} \right) = 21.7 \text{ g FeSAs}$$

Review And Reinforcement Stoichiometry Answers

Mark Jackson



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Microcomputer Software Directory, 1983 Community and Junior College Journal, 1976 *Whole Earth Software Catalog for 1986* Stewart Brand, 1985 **Ceramic Abstracts**, 1971 *Science Books & Films*, 1981 **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **Chemistry Equations and Answers** Mark Jackson, 2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students

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