

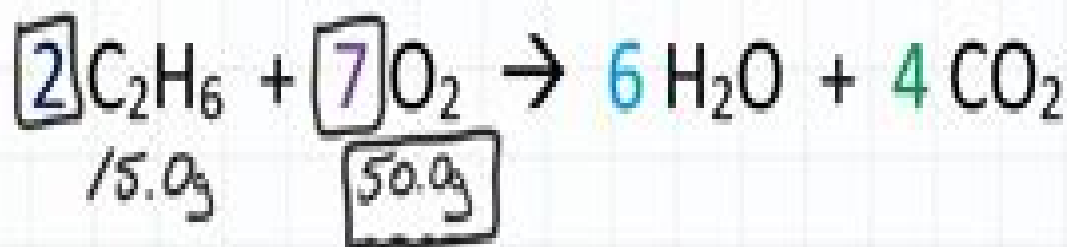
How to Identify the Limiting & Excess Reactants/Reagents

Molar Mass:

$$\text{C}_2\text{H}_6 = 30.08 \text{ g/mol}$$
$$O_2 = 32.00 \text{ g/mol}$$
$$\text{H}_2\text{O} = 18.02 \text{ g/mol}$$

1.) Calculate how much of one reactant will react with the other reactant

Ex. What is the limiting reactant when 15.0 g of C_2H_6 react with 50.0 g of O_2 ?



$$\frac{15.0 \text{ g } \text{C}_6\text{H}_6}{30.08 \text{ g } \text{C}_6\text{H}_6} \left(\frac{1 \text{ mol } \text{C}_6\text{H}_6}{1 \text{ mol } \text{C}_6\text{H}_6} \right) \left(\frac{7 \text{ mol } \text{O}_2}{2 \text{ mol } \text{C}_6\text{H}_6} \right) \left(\frac{32.00 \text{ g } \text{O}_2}{1 \text{ mol } \text{O}_2} \right) = 55.9 \text{ g } \text{O}_2$$

$$50.0 \text{ g } \overset{A}{\text{O}_2} \left(\frac{1 \text{ mol } \text{O}_2}{32.00 \text{ g } \text{O}_2} \right) \left(\frac{2 \text{ mol } \text{C}_2\text{H}_6}{7 \text{ mol } \text{O}_2} \right) \left(\frac{30.07 \text{ g}}{1 \text{ mol } \text{C}_2\text{H}_6} \right) = 13.4 \text{ g } \overset{B}{\text{C}_2\text{H}_6}$$

$$Q_2 = L.R.$$
$$C_2H_6 = E.R.$$

If reacting all of Reactant A requires **more** of Reactant B than is available, then **B** is the Limiting Reactant.

* If reacting all of Reactant A requires less of Reactant B than is available, then A is the Limiting Reactant.

Review Answers Limiting Reactants

Leo J. Malone, Theodore O. Dolter



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