



# Physical Science

## PDF Study Guide

**Properties of Elements**  
 Study Guide: The Periodic Table, Ions, and...  
 Questions and Answers  
 1. Q: What are the main characteristics of metals?  
 A: High electrical and thermal conductivity, and a shiny appearance.  
 2. Q: Where are non-metals located on the periodic table?  
 A: On the right side of the periodic table.  
 3. Q: What is a noble gas?  
 A: A gas that is chemically inert.  
 4. Q: Name a characteristic of noble gases.  
 A: They are inert and rarely form compounds.  
 5. Q: What type of elements are found in the middle of the periodic table?  
 A: Metals.  
 6. Q: What type of elements are found on the far left of the periodic table?  
 A: Alkali metals.  
 7. Q: What type of elements are found on the far right of the periodic table?  
 A: Noble gases.

**Properties of Elements**  
 Study Guide: The Periodic Table, Ions, and Valence Electrons  
 1. Overview of the Periodic Table  
 The periodic table is a systematic arrangement of elements based on their atomic number, electron configurations, and recurring chemical properties. Elements are organized into horizontal rows (periods) and vertical columns (groups). The periodic table is divided into three main categories: metals, non-metals, and metalloids (metals in between).  
 2. Metals, Non-Metals, and Metalloids  
 Metals are located on the left side of the periodic table. They are characterized by their ability to conduct electricity and heat, their malleability, and their shiny appearance. Examples include iron, copper, and aluminum.  
 Non-metals are located on the right side of the periodic table. They are generally poor conductors of electricity and heat, and they are brittle. Examples include carbon, oxygen, and nitrogen.  
 Metalloids are located in between the metals and non-metals. They have properties that are intermediate between metals and non-metals. Examples include silicon, boron, and arsenic.  
 3. Periodic Trends  
 The periodic table shows several trends in the properties of elements. These trends include atomic size, ionization energy, and electronegativity. Atomic size increases from top to bottom and from right to left. Ionization energy decreases from top to bottom and from right to left. Electronegativity increases from bottom-left to top-right.

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**Daniela Niemeyer**



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