

Parts of the Atom

**Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS
proton	proton	Nucleus	2000 X More than an e ⁻
neutron	neutron	Nucleus	2000 X More than an e ⁻
electron	electron	Orbitals / Energy Levels	1

**Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

6
C
carbon
12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element.

Error: Reference source not foundelement name: This is the name of the element.

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon:

Atomic mass: 12 (rounded)

Number of neutrons: $12 - 6 = 6$

To find the number of electrons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of electrons: 6

**Complete the table for the elements with atomic numbers 1-10.*

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

The Atom Cstephenmurray Answers

Jack Challoner



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The Structure of the Atom Stephen Miall, 1922 **The Question of the Atom** Mary J. Nye, 1983-01-01 **The Atom** Jack Challoner, 2018 Until now popular science has relegated the atom to a supporting role in defining the different chemical elements of the periodic table This bold new title places its subject center stage shining the spotlight directly onto the structure and properties of this tiniest amount of anything it is possible to identify The book covers a huge range of topics including the development of scientific thinking about the atom the basic structure of the atom how the interactions between atoms account for the familiar properties of everyday materials the power and mystery of the atomic nucleus and what the mysterious quantum realm of subatomic particles and their interactions can tell us about the very nature of reality Sparkling text banishes an outdated world of dull chemistry as it brightly introduces the reader to what everything is made of and how it all works on the most fundamental level **The Atom Uncovered** Mark Andrew Newstead, Stephen Charles Newstead, 2010 In this book we show how electromagnetic waves can form the sub atomic particles i e the proton neutron and electron as well as how these can fit together to form the nuclei of atoms and of all their isotopes From this we are able to combine the weak and strong force into the electromagnetic force *Atom (Icon Science)* Piers Bizony, 2017-05-04 Riddled with jealousy rivalry missed opportunities and moments of genius the history of the atom s discovery is as bizarre as capricious and as weird as the atom itself John Dalton gave us the first picture of the atom in the early 1800s Almost 100 years later the young misfit New Zealander Ernest Rutherford showed the atom consisted mostly of space and in doing so overturned centuries of classical science It was a brilliant Dane Neils Bohr who made the next great leap into the incredible world of quantum theory Yet he and a handful of other revolutionary young scientists weren t prepared for the shocks Nature had up her sleeve This insightful compelling book New Scientist reveals the mind bending discoveries that were destined to upset everything we thought we knew about reality and unleash a dangerous new force upon the world Even today as we peer deeper and deeper into the atom it throws back as many questions at us as answers *The Structure of the Atom* Stephen Miall, 1922 **Taming the Atom** Hans Christian Von Baeyer, 1995-10-01 Fascinating accessible study covers the components of the atom quantum mechanics the atomic landscape atoms in isolation quantum reality and more Lucid and entertaining The New York Times Book Review *The Atom* George Paget Thomson, 1962 **Questions and Answers on the Atom for Progress and Peace** , 1954 *The Study of the Atom* Francis Preston Venable, 1904 **Introducing the Atom** Roslyn D. Leeds, 1967 An introduction to nuclear physics through a definition of the atom and an analysis of its functions and behavior Includes a biographical listing of major scientists who have contributed to the field of nuclear physics as well as a glossary of terms **The Mysteries of the Atom** Harold Albert Wilson, 1934 **The Britannica Guide to the Atom** Britannica Educational Publishing, 2010-10-01 Common to all matter and its smallest and most fundamental unit the atom determines the properties of a substance and how it responds to other materials and stimuli Dividing the atom further

reveals several smaller particles whose activity forms the heart of nuclear physics Readers will discover the structure of the atom as well as the ways in which its electrons protons and neutrons facilitate both radioactivity and nuclear reactions both of which have been crucial to technological advancements since the 20th century

Questions and Answers on the Atom for Progress and Peace ,1954 The Project Physics Course Harvard Project Physics,F. James Rutherford,1975 The HISTORY of the ATOM the PERIODIC TABLE and RADIOACTIVITY Alex Kimmons,2021-05-15

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