

## EXPLORE: Visualizing Mixing with *Sugar and Salt Solutions* PhET simulation

### Learning Objectives:

- Describe the process of solution formation for salts versus molecular compounds.
- Identify the differences in an ionic solution and a molecular solution

**Directions:** Open the *Sugar and Salt Solutions* simulation [CLICK HERE](#)

**Move the conductivity tester to the solution.  
Shake some salt into the water.**

1. Does the conductivity tester light up? **Yes.**
2. What kind of substance causes the conductivity tester to light? **Electrolyte.**

**Click on the "remove salt" button.  
Change the solute to sugar and shake some into the water.**

1. Did it light up? **No.**
2. What type of substance is this? **Nonelectrolyte.**

**Click on the "Micro" tab at the top of the page.  
Select the sodium chloride solute.  
Shake some salt into the water.**

1. What happens to the salt as it mixes with the water? **The salt ions freely move throughout the water and the salt concentration increases.**

**Look at the Concentration graphic.**

2. How do the concentrations of the 2 ions compare?  
**The salt and chloride ions equally concentrate the water.**

**Now move the evaporation slider to the right and hold until all the water has evaporated.**

1. What happened to the concentration of the ions as the water evaporates? **The concentration of the ions increases as the water evaporates.**
2. What do you observe? **The ions also don't move as freely**
3. Is the substance, NaCl, different than when it was first added to the water? **Yes, the substance joins into one compound instead of two.**

**Click the reset button at the bottom.  
Select the solute glucose.  
Add the sugar to the water**

1. What do you observe? **I observe that the Glucose molecules remain intact, they don't break apart into separate atoms in the water.**
2. How does the process of solvation for sugar differ from that of salt?  
**The process of solvation for sugar allows the sugar molecules to stay intact in the water, while the process of solvation for salt breaks the compound apart into ions.**
3. Which solution results in a greater number of particles in solution- salt or sugar? **Salt**

**Use the slider at the bottom to evaporate the water.**

1. What did you observe? How is the formation of sugar crystals different from the formation of salt crystals?  
**I observed that both the sugar and the salt particles joined together as one compound when the water was evaporated. The formation of sugar crystals consists of sugar molecules that stay**

# Vbpf Simulation Worksheet Key

**B Lingard**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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