

VCE Chemistry Unit 3 Area of Study 2

How can the yield of a chemical product be optimised?

Rate of Chemical Reactions

- Chemical Reactions with reference to the Collision Theory
 - Qualitative interpretation of Maxwell-Boltzmann Distribution Curves
- Exothermic vs Endothermic Reactions
 - Enthalpy Changes
 - Representations in Energy Profile Diagrams
- Factors Affecting Rate of Chemical Reactions
 - Temperature, Surface Area of Solids, Concentration of Solutions, Gas Pressure, Presence of a Catalyst
- Role of Catalysts in changing the rate of Chemical Reactions

Extent of Chemical Reactions

- Reversible vs Irreversible Reactions
- Rate vs Extent of Reactions
- Heterogeneous Equilibria
 - Reference to Collision Theory and Concentration-Time Graphs
- Calculations Involving Equilibrium Expressions and Equilibrium Constants for a closed homogeneous equilibrium system
- Le Chatelier's Principle
 - Factors that favour the yield of a chemical reaction
 - Representation of Equilibrium system changes using Concentration-Time Graphs
 - Combusting Equilibria → Carbon Monoxide Poisoning

Production of Chemicals by Electrolysis

- Electrolysis of molten liquids and aqueous solutions using different electrodes
- Operating principles of commercial electrolytic cells
 - Selection of a suitable electrolyte (molten or aqueous)
 - Selection of suitable electrodes (active or inert)
- Use of the Electrochemical Series to explain or predict products of electrolysis
- Electrolytic Cells vs Galvanic Cells
 - Energy Transformations, Basic Structural Features and Processes
- Faraday's Laws and Stoichiometry
 - Determining amounts of product, current or time for an electrolytic process

Rechargeable Batteries

- Rechargeable Batteries (secondary cells)
 - Discharge as a Galvanic Cell
 - Recharge as an Electrolytic Cell
- Factors affecting Battery Life
 - Components and Temperature

Vcaa Chemistry Study Guide

Sebastian Brünink



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Fortify VCE Chemistry Units 3 and 4 2017-2021 Study Guide Joseph Rowan,Sophie Zhou,Emma Zheng,2020-07
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