

1. The Structure of the Atom

- The Particle Model states that all matter is made up of atoms.
- Atoms have origin in stars.
- We have been able to modify theories and develop new ones.

Nanomaterials

- Have diameters of 1-100 nm - nanometre as a useful unit.
- May have different properties than bulk materials.
- Large SA compared to V - surfaces are used to absorb and transport molecules as well as being catalysts for chemical reactions
- Measurements should be converted as below:

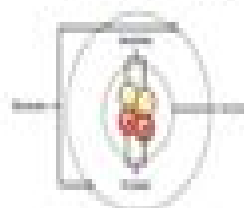
$$1 \text{ m} = 10^3 \text{ cm} = 10^3 \text{ mm} = 10^6 \mu\text{m} = 10^9 \text{ nm}$$

Historical development

1808	Dalton - elements have very small particles - atoms
1897	Thomson - positively charged spheres with negatively charged particles - electrons
1911	Rutherford - protons in the centre, electrons surrounding, mass concentrated in the nucleus
1913	Bohr - electrons orbit the nucleus in different energy levels, changing orbits - emit/gain energy
1926	Schrodinger - major electrons energy levels - subshells, orbitals are grouped together in the same subshell
1932	Chadwick - discovered the nucleus, explained the existence of isotopes

Location of particles

- Positively charged protons and uncharged neutrons have the same mass, both contribute the most to the atomic mass
- Negatively charged electrons move around the nucleus in specific energy levels, mass is apx. 1800 times smaller



Representing elements

- Atomic number (symbol Z) is the number of protons in the atom
- Mass number (symbol A) is the number of protons **and** neutrons
- In a neutral atom protons = electrons
- Mass number - Atomic number = number of neutrons
- Representing elements and isotopes is as below:



Isotopes

- Isotopes have same atomic number but different mass number
- isotopes have same chemical properties but different physical properties (mass, radioactivity, density)



Vce Chemistry Aspirin Sac

Victor M. Corman



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