

# STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

### STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass.
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each.

### Using Ratios

Reactant and product amounts are related by ratios.

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction. For the next example, let's use only moles of chloride.

Consider the reaction:

$$2\text{NaCl} + \text{MgCl}_2 \rightarrow 2\text{NaCl}_2 + \text{MgCl}_2$$

We can write these ratios:

moles of sodium	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$
moles of chlorine	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$
moles of oxygen	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$
moles of hydrogen	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$
moles of carbon	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$
moles of sulfur	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}_2}{2\text{NaCl}_2}$

### Example: Finding the Limiting Reactant

In the reaction:

$$2\text{NaCl}_2 + \text{MgCl}_2 \rightarrow 2\text{NaCl}_2 + \text{MgCl}_2$$

5.37 grams of  $\text{NaCl}_2$  react with 0.43 grams of  $\text{Mg}$ . What is the limiting reactant and what mass of  $\text{MgCl}_2$  can be formed?

First, we need to write the balanced equation:

$$2\text{NaCl}_2 + \text{MgCl}_2 \rightarrow 2\text{NaCl}_2 + \text{MgCl}_2$$

Next, we need to write the ratios:

$$\frac{2\text{NaCl}_2}{2\text{NaCl}_2} = \frac{1\text{MgCl}_2}{1\text{MgCl}_2}$$

Then, we need to write the ratios:

$$\frac{5.37\text{g NaCl}_2}{58.44\text{g/mol}} = \frac{0.0919\text{mol NaCl}_2}{1}$$
$$\frac{0.43\text{g Mg}}{24.31\text{g/mol}} = \frac{0.0177\text{mol Mg}}{1}$$

Now, we need to compare the ratios:

$$\frac{0.0919\text{mol NaCl}_2}{0.0177\text{mol Mg}} = \frac{5.19}{1}$$

Since the ratio of  $\text{NaCl}_2$  to  $\text{Mg}$  is greater than the ratio of  $\text{NaCl}_2$  to  $\text{Mg}$  in the balanced equation,  $\text{Mg}$  is the limiting reactant.

### Using Mole Ratios

From the last page, we found that the reaction:

$$2\text{NaCl}_2 + \text{MgCl}_2 \rightarrow 2\text{NaCl}_2 + \text{MgCl}_2$$

has the following ratios of moles of each substance using the coefficients of the balanced equation:

Now, let's use the ratios we found:

How many moles of  $\text{Mg}$  can be produced from 5.37 moles of  $\text{NaCl}_2$ ?

We know the ratio of moles of  $\text{Mg}$  to moles of  $\text{NaCl}_2$  is 1:2, so we can write the ratio as  $\frac{1\text{Mg}}{2\text{NaCl}_2}$ . By writing this ratio with the denominator, we can use the ratio to find the moles of  $\text{Mg}$ .

$$5.37\text{mol NaCl}_2 \times \frac{1\text{Mg}}{2\text{NaCl}_2} = 2.69\text{mol Mg}$$

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