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web dec 25 2013 modified 9 years 8 months ago viewed 2k times 1 i want to design a combinational circuit for a vending machine that delivers 2 items tea 2 and water 1 the machine accepts up to 3 dollars i draw the truth table with 3 inputs dollars and 2 outputs tea water

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web jan 5 2015 1 hi everyone i am currently studying electronics at college and have been tasked with designing a circuit to do the following a hot drink vending machine can give either tea or coffee with or without milk with or without sugar

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web 1 consider the reaction  $\text{i}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$  a 80.0 grams of iodine pentoxide  $\text{i}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $\text{I}_2$  was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web chem 12.3 limiting reagent and percent yield 4.06 reviews whenever quantities of two or more reactants are given in a stoichiometric problem you must identify the limiting reagent

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $0.216 \text{ mol } \text{O}_2 \times \frac{3 \text{ mol CO}}{3 \text{ mol } \text{O}_2} = 0.216 \text{ mol CO}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 160.0 \text{ g } \text{Al}_2\text{O}_3$  c determine the percent yield  $\frac{160.0 \text{ g}}{160.0 \text{ g}} \times 100 = 100\%$

**7 2 limiting reagent and reaction yields chem 1114** - Jul 02 2022

web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web limiting reagents and percentage yield worksheet 1 consider the reaction i<sub>2</sub>o<sub>5</sub> g 5 co g 5 co<sub>2</sub> g i<sub>2</sub> g a 80 0 grams of iodine v oxide i<sub>2</sub>o<sub>5</sub> reacts with 28 0 grams of carbon monoxide co determine the mass of iodine i<sub>2</sub> which could be produced

80 g i2o5 1 mol i2o5 1 mol i2 xs 1 333 8 g i2o5 1 mol i2o5 28 g co 1 mol co

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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