

- When an oxygen atom forms an oxide ion,  $O^{2-}$ , it achieves a noble gas configuration.
  - What is a noble gas configuration?  
*It fills outer level*
  - What noble gas does oxide ion,  $O^{2-}$ , have the electron configuration of?  
*Neon*
- When an aluminum atom forms an aluminum ion,  $Al^{3+}$ , it becomes more stable.
  - Are electrons lost or gained during this change?  
*lost*
  - How many electrons are lost/gained?  
*Three electrons are lost.*
  - Why is an aluminum ion,  $Al^{3+}$ , more stable than an aluminum atom?  
 *$Al = 1s^2 2s^2 2p^6 3s^2 3p^1$ .  $Al^{3+}$  has none. It achieves an  $1s^2 2s^2 2p^6$  have the same electron arrangement as neon giving it a full outer level which is more stable.*
- Why do most transition elements exhibit more than one possible oxidation number?  
*Because they can lose electrons from different energy levels and do not have to lose them all at once but instead can lose some first and so on.*
- Nearly all transition elements can have oxidation numbers of  $\pm 1$ . Explain why?  
*More than 2 valence electrons have electrons which they can lose.*
- The electron configuration of iron, Fe, is  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$ .
  - Write the electron configuration for  $Fe^{2+}$ .  
 *$1s^2 2s^2 2p^6 3s^2 3p^4$*
  - Write the electron configuration for  $Fe^{3+}$ .  
 *$1s^2 2s^2 2p^6 3s^2 3p^4$*
  - Explain why  $Fe^{2+}$  ion tends to form.  
*This has a 1/2 full outer shell which is an stable as it can get closer to nearest noble gas electrons and the  $Fe^{2+}$ .*

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