

**MATH SKILLS TRANSPARENCY WORKSHEET****6****Using the Periodic Table****Use with Chapter 6,  
Section 6.2**

1. Identify the number of valence electrons in each of the following elements.

a. Ne \_\_\_\_\_ e. O \_\_\_\_\_  
b. K \_\_\_\_\_ f. Cl \_\_\_\_\_  
c. B \_\_\_\_\_ g. P \_\_\_\_\_  
d. Mg \_\_\_\_\_ h. Si \_\_\_\_\_

2. Identify the energy level of the valence electrons in each of the following elements.

a. Br \_\_\_\_\_  
b. N \_\_\_\_\_  
c. Ra \_\_\_\_\_  
d. H \_\_\_\_\_  
e. Ar \_\_\_\_\_  
f. I \_\_\_\_\_

3. Use the periodic table to write the electron configurations (using noble gas notation) for each of the following elements.

a. Li \_\_\_\_\_  
b. F \_\_\_\_\_  
c. As \_\_\_\_\_  
d. Sr \_\_\_\_\_  
e. Bi \_\_\_\_\_

4. Determine the group, period, and block of the elements having the following electron configurations.

a.  $1s^2$  \_\_\_\_\_  
b.  $[\text{Ne}]3s^23p^1$  \_\_\_\_\_  
c.  $[\text{Ar}]4s^1$  \_\_\_\_\_  
d.  $[\text{Kr}]5s^24d^1$  \_\_\_\_\_  
e.  $[\text{Xe}]6s^24f^{14}5d^{10}6p^4$  \_\_\_\_\_

# Teaching Transparency The Periodic Table

**Rosina Ehmann**



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