- The electron configuration of hydrogen is like that of Group 1 metals.
- Magnesium is more reactive than iron and will replace iron in a compound. If the iron corrodes, the magnesium will replace it, restoring the iron.
- The number of oxygen atoms is not balanced because parentheses were not placed around the hydroxide ion.

Applying Scientific Methods

- 2CuS(s) + 3O₂(g) → 2CuO(s) + 2SO₂(g);
 2Cu₂S(s) + 3O₂(g) → 2Cu₂O(s) + 2SO₂(g)
- 2. 2.5, 4, 4, 2, 2
- a. CuO(s) + C(s) → Cu(l) + CO(g); Cu₂O(s) + C(s) → 2Cu(l) + CO(g)
 - b. Sample answer: Mg(l) + CuO(s) → Cu(l) + MgO(s)
- 2Fe₂O₃(s) + 3C(s) → 4Fe(l) + 3CO₂(g)
- 5. $4Fe(s) + 3O_2(g) \rightarrow 2Fe_2O_3(s)$
- 6. synthesis.
- Na(I) + KCI(I) → NaCl(I) + K(g)
- 8. $Ca(1) + 2CsCl(1) \rightarrow CaCl_2(1) + 2Cs(g)$

CHAPTER 10

MiniLab 10 - Percent Composition and Gum

Analysis

- Mass dissolved = (initial mass) (final mass) = 3.11 g - 2.84 g = 0.27 g dissolved sweeteners and flavorings
- Mass dissolved = (initial mass) (final mass) = 3.11 g - 2.75 g = 0.36 g dissolved sweeteners and flavorings
- whole piece of gum: 0.27 g/3.11 g × 100 = 8.7% small pieces of gum: 0.36 g/3.11 g × 100 = 12%
- Because more sweetener dissolves when the surface area is larger, the inference is that the sweeteners are mixed throughout the gum.

Expected Results: The gum will lose mass after being placed in water because the sweeteners are water-soluble. The cut piece with the larger surface area will expose more of the soluble chemicals to the water, and will have a greater loss of mass.

ChemLab 10 - Hydrated Crystals

Pre-Lab

- To obtain the mass of water, subtract the mass of the crucible, cover, and anhydrous MgSO₄ from the mass of the crucible, cover, and hydrated MgSO₄. To obtain the mass of anhydrous MgSO₄, subtract the mass of the crucible and cover from the mass of the crucible, cover, and anhydrous MgSO4.
- 4. Divide by the molar mass.
- Calculate the ratio of moles of water to moles of anhydrous MgSO₄.

Mass Data and Observations of Epsom Salts	
Observations of hydrated MgSO _a	Shiny. translucent crystals
Mass of crucible and lid	11.36 g
Mass of crucible, list, and hydrated MgSO ₄	14.36 g
Mass of hydrated MgSO ₄	3.00 g
Mass of crucible, list, and anhydrous MgSO ₄	12.83 g
Main of anhydrous MgSO,	1.47 g
Mass of water in hydrated MgSO _a	1.53 g
Moles of anhydrous MgSO ₄	0.0122 mol
Moles of water in hydrated MgSO _a	0.0849 mol
Observation of anhydrous MgSO ₄	Bright white, opaque, powdery

Analyze and Conclude

- MgSO₄-7H2O.
- The hydrated MgSO₄ is shiny and translucent. The anhydrous MgSO₄ is opaque and bright white.

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