1.2.1 Compare (qualitatively) the relative strengths of ionic, covalent, and metallic bonds.

1. Compare the relative strengths of the three different types of bonds.

ionic > covalent > metallic

2. Which statement compares the amount of energy needed to break the bonds in $CaCl_2$ (E1) and $C_{12}H_{22}O_{11}$ (E2)?
a. E1>E2, as $CaCl_2$ is a covalent compound.
b. E1<E2, as $CaCl_2$ is a covalent compound.
(c) E1>E2, as $CaCl_2$ is an ionic compound.
d. E1<E2, as $CaCl_2$ is an ionic compound.

1.2.2 Infer the type of bond and chemical formula formed between atoms.

1. Which pair of elements would ***most likely*** bond to form a covalently bonded compound?
a. sodium and fluorine
b. barium and chlorine
(c) phosphorus and oxygen
d. magnesium and sulfur

2. For each pair of atoms, predict whether the bond formed between the atoms is either ionic or covalent, and write the formula for t
predicted compound.

ionic = M+NM (or polyatomic ion)
Covalent = all NM
Metallic = all M

1. Na and O ionic, Na_2O
2. S and F Covalent, SF_2
3. Ag and N ionic, Ag_3N
4. Te and H covalent, TeH_2

3. Which statement describes the compound formed between sodium and oxygen?
a. It is NaO_2, which is ionic.
b. It is NaO_2, which is covalent.
(c) It is Na_2O, which is ionic.
d. It is Na_2O, which is covalent.

1.2.3 Compare inter- and intra- particle forces.

1. Rank the following substances in the order in which they would evaporate, justifying the order of placement for each (using those
inter- and intra- particle forces).
1. Water (H_2O)
2. Methane (CH_4)
3. Sodium chloride (NaCl)
4. Phosphorus trifluoride (PF_3)

1. CH_4 - LDFs (weakest)
2. PF_3 - LDFs, dipole-dipole
3. H_2O - LDFs, dipole-dipole, H-bonding
4. NaCl - ionic bond (strongest force)

2. At STP, fluorine is a gas and iodine is a solid. Why?
a. Fluorine has lower average kinetic energy than iodine.
b. Fluorine has higher average kinetic energy than iodine.
(c) Fluorine has weaker intermolecular forces of attraction than iodine.
d. Fluorine has stronger intermolecular forces of attraction than iodine.

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