

# AP World History Review Guide

Especially in treatment studies, how they both treated patients, including a control group, is a crucial factor in how results are generally interpreted. It is, as mentioned, all too often ignored, perhaps due to laziness. When practitioners themselves used very different methods, though, no matter whether one group is better, most groups of independent studies (up to 1990) were like this. However, most today do recognize that a low number of trials that do not provide significant support for a particular treatment is insufficient to rule out effectiveness, although few recognize that these independent studies often will be used to claim effectiveness, even when they do not achieve a statistically significant result. In fact, some people will say that the question is whether independent studies are like other treatments, or whether a specific drug is exceptional. But, given the data, why not think about this other way around, so that there are very few studies that are better than other ones, even though some may not appear to be statistically significant?

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- \* \* S — Special
- \* \* P — Political
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# World History Review Guide

**James Zucker**



## World History Review Guide:

ASAP World History: Modern, 2nd Edition: a Quick-Review Study Guide for the AP Exam The Princeton Review, 2019

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*Study Guide for World History I* United States Armed Forces Institute, 1956

*AP World History Study Guide 2019* Inc. AP Exam Prep Team Accepted, 2018-07-26

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although the c-cl bonds are rather polar the individual bond dipoles cancel one another in this symmetrical structure and does not have a net dipole moment

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 determine based on electronegativity and symmetry if the following molecules are polar or nonpolar  
 hcn h<sub>2</sub>c n<sub>2</sub> cf<sub>4</sub> c<sub>2</sub>f<sub>4</sub> co<sub>2</sub> cl<sub>2</sub> nf<sub>3</sub> n<sub>2</sub>f<sub>3</sub> q3

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web when the atoms in a bond are not the same the bonding electrons are shared and the bond is electronegativities equally unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by consulting a table of

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 a. cl<sub>2</sub> b. clo<sub>2</sub> c. tecl<sub>2</sub> d. pcl<sub>3</sub> e. sef<sub>4</sub> f. ph<sub>2</sub> g. xef<sub>2</sub> answer problem 6.2.3.6.2.3

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web the relative electronegativities of the bonded atoms is h < c < cl and so the bond moments all point toward the cl end of the

molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as  $\text{BF}_3$  trigonal planar  $\text{CH}_4$  tetrahedral  $\text{PF}_5$  trigonal bipyramidal and  $\text{SF}_6$

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