

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride  $\text{SF}_6$
- Lithium nitride  $\text{Li}_3\text{N}$
- Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride  $\text{SnCl}_2$
- Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- Potassium bromate  $\text{KBrO}_3$
- Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- Chloric acid  $\text{HClO}_3$
- Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds:

- $\text{CuSO}_4$  Copper (II) sulfate
- $\text{AlF}_3$  aluminum fluoride
- $\text{HI}$  hydrogen iodide
- $\text{NO}$  nitrogen monoxide
- $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- $\text{HNO}_3$  nitric acid
- $\text{HNO}_2$  nitrous acid
- $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

# World Of Chemistry Study Guide Answers

**Neil E. Schore**



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