

Titration calculations

In a titration HCl was titrated against 25cm³ of a standard solution of KOH (0.20 mol/dm³). 28.5cm³ of HCl was needed to completely neutralise the KOH.

Calculate the concentration of the HCl.

	HCl	+	KOH	□	KCl	+	H ₂ O
Amount (mol)							
Volume (cm ³)							
Volume (dm ³)							
Conc (mol/dm ³)							

Step 1

Write in the table what values we know already from the question (Concentrations/volumes)

Step 2

Convert volumes into dm³ (÷ 1000)

Step 3

Work out the number of moles of KOH we have (Moles = conc x volume)

Step 4

Look at the ratio to work out the number of moles of HCL

Step 5

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**Daniel W Foster Professor of Medical
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