

- 1) It takes 83 mL of a 0.45 M NaOH solution to neutralize 235 mL of an HCl solution. What is the concentration of the HCl solution?

- 2) You are titrating an acid into a base to determine the concentration of the base. The endpoint of the neutralization is reached but the stopcock on the buret sticks slightly and allows a few more drops of acid to fall into the solution. How will this affect your calculations for the concentration of the base?

- 3) It takes 38 mL of 0.75 M NaOH solution to completely neutralize 155 mL of a sulfuric acid solution (H_2SO_4). What is the concentration of the H_2SO_4 solution?

- 4) A few small drops of water are left in a buret that is then used to titrate a base into an acid solution to determine the concentration of the acid. Will this small amount of water have any effect on the determined value for the concentration of the acid? If so, how is it affected?

- 5) It takes 12.5 mL of a 0.30 M HCl solution to neutralize 285 mL of NaOH solution. What is the concentration of the NaOH solution?

- 6) Lulu Labwrecker carefully pipets 25.0 mL of 0.525 M NaOH into a test tube. She places the test tube into a small beaker to keep it from spilling and then pipets 75.0 mL of 0.355 M HCl into another test tube. When Lulu reaches to put this test tube of acid into the beaker along with test tube of base she accidentally knocks the test tubes together hard enough to break them and their respective contents combine in the bottom of the beaker. Is the solution formed from the contents of the two test tubes acidic or basic? What is the pH of the resulting solution?

Titration Worksheet Chemfiesta

**Izaak Maurits Kolthoff, Nathaniel
Howell Furman**



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Volumetric Analysis Ralf Martens-Menzel, Lena Harwardt, Hanns-Jürgen Krauss, 2025-05-19 This standard work on volumetric analysis based on the 20th German edition provides comprehensive information on the theory of acid base titration redox titration complexation titration and precipitation titration with both classical and instrumental indication of the equivalence point Many applications are described and explained in detail with examples in pharmaceutical and environmental analysis

Complexometric Titrations Gerold Schwarzenbach, 1957

Working with Chem Separate Wink, 1999-12-01

EDTA Titrations H. A. Flaschka, 2013-10-22 EDTA Titrations An Introduction to Theory and Practice Second Edition considers the theoretical background full procedural details and some practical applications of EDTA titrations Ethylenediamine tetra acetic acid EDTA has risen from an obscure chemical compound to the most widely used organic reagent This book is composed of 21 chapters The opening chapters present the general theoretical foundations of EDTA titrations The subsequent chapters describe the properties of EDTA such as the stability constants titration curves selectivity and masking effect These topics are followed by discussions on titration types standard solutions and reagents The remaining chapters cover some of the practical applications of EDTA titrations This book is directed toward students with advanced courses in analytical and organic chemistry

Titration in Non-aqueous Media István Gyenes, 1968

Advanced Potentiometry Erzsébet Néher-Neumann, 2010-03-14 Advanced Potentiometry details the accurate calculation of potentiometric titrations Additional terms such as the ideal diffusion potential together with the contribution of activity factors furnish calculated stability constants which do not include systematic errors and thus present a complete and correct description of equilibrium processes Moreover serious systematic errors resulting from the recent use of curve fitting computer programs are highlighted later in the text Advanced Potentiometry can advantageously be used by lecturers researchers PhD graduates and undergraduate students working in the fields of Solution Chemistry Physical Chemistry Inorganic Chemistry and Analytical Chemistry and Environmental Research

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A Teacher's Guide on Complexometric Titration Shoukat Ali R A,2019-12-13 Lesson Plan from the year 2019 in the subject Chemistry Anorganic Chemistry grade A language English abstract This laboratory material seeks to help fresh teachers of the UG PG departments as well as self studying students It eliminates the difficulties which are common in the starting stage of a teaching carrier regarding solution preparations in various concentrations calculations and procedures for the experiments and the practical set up Moreover the work is helpful to understand the role of reagents chemicals used in experiments reactions conditions and structures The work will provide all the information related to the complexometric titrations

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