- 1. How many electrons are in the following elements?
- A. Carbon.

C. Indium

E. Tungsten.

B. Chlorine

D. Calcium

F. Antimony

2. Draw Lewis Structures for each of the following elements. How many electrons are in their outer shells?

A. Silicon

C. Sulfiur

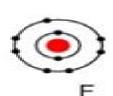
E. Bervillium

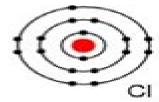
B. Potassium.

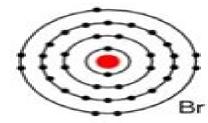
D. Boron

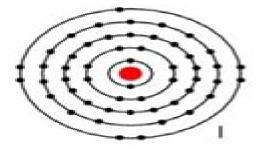
F. Argon

3. Circle the valence electrons only. How many valence electrons does each atom have?









4. Finish the sentence: Elements in the same group have the same number of electrons.

5. Fill in the electrons for the following elements. Use your work to answer how many valence electrons each element has.

Example: C (carbon): 1ST+ 2ST+ 2PT T 3S 3P #valence electrons 4

He: 1S 2S 2P 3S 3P ____

Number of valence electrons in helium _____

Cl: 1S 2S 2P 3S 3P

Number of valence electrons in chlorine _____

Li: 1S 2S 2P 3S 3P B: 1S 2S 2P 3S 3P

Number of valence electrons in lithium _____

Number of valence electrons in boron _____

F: 1S __2S __2P ______3S __3P ______

Number of valence electrons in fluorine

Ar: 1S 2S 2P 3S 3P

Number of valence electrons in argon _____

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