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so calculating enthalpy change of reaction in 2 $\Delta H_{\text{reaction}} = 6 \times \Delta H_{\text{f}}^\circ \text{ CH}_4\text{ (g)}$ $\Delta H_{\text{f}}^\circ \text{ C}_2\text{H}_2\text{ (g)}$ $6 \times 523 - 82 - 3140 = 82$

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