

## SECTION-II

### LOGIC OF INORGANIC REACTIONS

#### TYPES OF INORGANIC REACTIONS

All inorganic reactions fall into any one of the following two categories.

(i) Metathesis (ii) Redox

#### (A) METATHESIS:

The reaction in which there is no change of *Oxidation Number (ON)* of any element is called a metathesis reaction. Some common metathesis types of reactions are discussed below.

##### (i) **Double Replacement (Ion Exchange) Reaction:**

This involves the reaction between two substances (which are mostly soluble in water) in which the basic and acid parts (radicals) of two substances get exchanged. That means the basic part of one substance joins with the acid part of the other substance and vice versa.

Double Replacement reaction is again of two types.

(a) neutralisation (acid-base reaction) reaction

(b) precipitation reaction

**Neutralisation:** An acid reacts with a base (alkali) to produce a salt and water. The sour taste of acid and bitter (soda like) taste of alkali kill each other and a salty taste is found in the product.



In all these cases you find that the ON of none of the elements has undergone any change. So these are metathesis reactions.

**IMPORTANT:** Note that in this chapter, you do not have to balance any equation. Our purpose is to know what products are formed during a reaction and why they are formed. In other words, we have to develop the logic of the reaction as to why a particular set of products are formed and why not any other products are formed. So you are categorically advised not to balance the equations throughout this chapter.

**Precipitation:** When two soluble substances undergo double replacement reaction and one of the products is insoluble and appears as a precipitate, it is called precipitation reaction. Look to this example.



Here in this double replacement reaction, AgCl appears as a curly white precipitate. Therefore it is called a precipitation reaction.

Let us take some more examples.



There are however some reactions which belong both to neutralisation as well as precipitation. See this example.

## Section 2neutralization Reactions

**Michal Rosen-Zvi**



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