

52**ELECTRON ARRANGEMENT IN ATOMS****Section Review****Objectives**

- Describe how to write the electron configuration for an atom
- Explain why the actual electron configurations for some elements differ from those predicted by the Aufbau principle

Vocabulary

- electron configurations
- Aufbau principle
- Pauli exclusion principle
- Hund's rule

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The ways in which electrons are arranged around the nuclei of atoms are called ____ 1 _____. The ____ 2 _____ describes the sequence in which orbitals are filled. The various orbitals within a sublevel of a principle energy level are always of ____ 3 _____ energy. The ____ 4 _____ principle states that a maximum of only ____ 5 _____ electrons can occupy each orbital. To occupy the same orbital, two electrons must have ____ 6 _____ spins. Hund's rule states that the electrons pair up only after each orbital in a sublevel is occupied by ____ 7 _____. When using the shorthand method for showing the electron configuration of an atom, ____ 8 _____ are used to indicate the number of ____ 9 _____ occupying each sublevel.

Correct electron configurations can be obtained by using the Aufbau diagram for the elements up to and including vanadium. ____ 10 _____ and copper are exceptions to the Aufbau principle.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Section 5 2 Electron Arrangement In Atoms Worksheet Answers

Padhraic Smyth



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