



12. Number of atoms in 560 g of Fe (atomic mass 56  $\text{g mol}^{-1}$ ) is :  
 (1) Twice that in 70 g N (2) Half that in 20 g H  
 (3) Both (1) and (2) (4) None of these
13. Which has maximum number of atoms :  
 (1) 24 g of C (12) (2) 56 g of Fe (56) (3) 27 g of Al (27) (4) 108 g Ag (108)
14. The total number of protons, electrons and neutrons in 12 g of  $^{12}_6\text{C}$  is :  
 (1)  $1.084 \times 10^{25}$  (2)  $6.022 \times 10^{23}$  (3)  $6.022 \times 10^{22}$  (4) 18
15. 1.0 g of hydrogen contains  $6 \times 10^{23}$  atoms. The atomic weight of helium is 4. It follows that the number of atoms in 1 g of He is :  
 (1)  $\frac{1}{4} \times 6 \times 10^{23}$  (2)  $4 \times 6 \times 10^{23}$  (3)  $6 \times 10^{23}$  (4)  $12 \times 10^{23}$
16. A sample of aluminium has a mass of 54.0 g. What is the mass of the same number of magnesium atoms? (At. wt. Al = 27, Mg = 24)  
 (1) 12 g (2) 24 g (3) 48 g (4) 96 g.
17. The number of atoms in 558.5 g of Fe (at wt. = 55.85) is :  
 (1) Twice that in 60 g carbon (2)  $6.022 \times 10^{23}$   
 (3) Half in 8 g He (4)  $558.5 \times 6.023 \times 10^{23}$
18. The total number of electrons present in 8.0 g of methane is.  
 (1)  $4.8 \times 10^{24}$  (2)  $3.01 \times 10^{24}$  (3)  $4.8 \times 10^{23}$  (4)  $3.01 \times 10^{23}$ .
19. How many moles of electron weigh one kilogram :  
 (1)  $6.023 \times 10^{23}$  (2)  $\frac{1}{9.108} \times 10^{31}$  (3)  $\frac{6.023}{9.108} \times 10^{24}$  (4)  $\frac{1}{9.108 \times 6.023} \times 10^6$
20. Under the same conditions, two gases have the same number of molecules. They must  
 (1) be noble gases (2) have equal volumes  
 (3) have a volume of 22.4  $\text{dm}^3$  each (4) have an equal number of atoms
21. 5.6 litre of oxygen at STP contains :  
 (1)  $6.02 \times 10^{23}$  atoms (2)  $3.01 \times 10^{23}$  atoms  
 (3)  $1.505 \times 10^{23}$  atoms (4)  $0.7525 \times 10^{23}$  atoms
22. If V ml of the vapours of substance at NTP weight W g. Then molecular weight of substance is :  
 (1)  $(W/V) \times 22400$  (2)  $\frac{V}{W} \times 22.4$  (3)  $(W-V) \times 22400$  (4)  $\frac{W \times 1}{V \times 22400}$
23. If 1  $\text{dm}^3$  of a gas weights 2.5 g at STP, its gram-molecular weight is :  
 (1) 56 g (2) 11.2 g (3) 22.4 g (4) 224 g
24. 16 g of an ideal gas  $\text{SO}_x$  occupies 5.6 L, at STP. The value of x is  
 (1)  $x = 3$  (2)  $x = 2$  (3)  $x = 4$  (4) none of these

### Section (C) : Density and Vapour density

1. Vapour density of a gas if its density is 0.178 g/L at NTP is :  
 (1) 0.178 (2) 2 (3) 4 (4) 0.089
2. A gas is found to have the formula  $(\text{CO})_x$ . It's VD is 70 the value of x must be :  
 (1) 7 (2) 4 (3) 5 (4) 6
3. Number of electrons in 1.8 mL of  $\text{H}_2\text{O}(l)$  is about :  
 (1)  $6.02 \times 10^{23}$  (2)  $3.011 \times 10^{23}$  (3)  $0.6022 \times 10^{21}$  (4)  $60.22 \times 10^{20}$

### Section (D) : Percentage composition and Molecular formula

1. The percentage value of nitrogen in urea is about  
 (1) 46 (2) 85 (3) 18 (4) 28

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**R Pring**



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