


How Substances Dissolve

KEY IDEAS

As you read this section, keep these questions in mind:


- Why is water called the universal solvent?
- How do substances dissolve?

Why Do Substances Dissolve in Certain Solvents?

Water is often called a universal solvent because many substances dissolve in it. However, no one substance can dissolve every solute. A general rule in chemistry is that “like dissolves like.” This rule means that a solvent will dissolve substances that have similar molecular structures. 

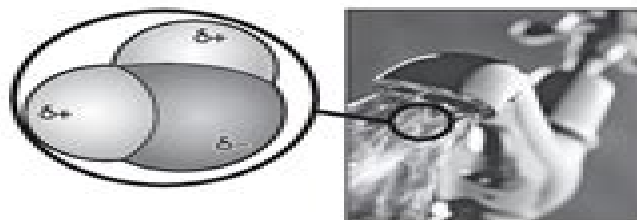
A solvent must be able to attract solute particles and pull them away from one another. Thus, solvent particles must attract solute particles more strongly than the solute particles attract one another.

POLAR MOLECULES

Recall that atoms bonded together do not all share electrons equally. For example, a water molecule is made up of two hydrogen atoms bonded to one oxygen atom. The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom. This gives each hydrogen atom a partial positive charge (δ^+) and the oxygen atom a partial negative charge (δ^-). 

Molecules that have partially charged positive and negative areas are **polar**. Because water molecules are polar, water can dissolve many other polar substances. Water can also dissolve ionic compounds such as table salt.

Water is a polar molecule because the oxygen atom strongly attracts electrons. Because the electrons spend more time near the oxygen atom, the oxygen atom has a partial negative charge.



READING TOOLBOX

Summarize Read each page of the section silently to yourself. With a partner, take turns summarizing each page.

READING CHECK

1. Identify Why is water called a universal solvent?

because
many substances dissolve
in it

READING CHECK

2. Explain Why does a water molecule have partial charges?

The oxygen atom attracts
electrons more strongly
than the hydrogen atoms
do. Thus, the electrons
spend more time near the
oxygen atom.

Section How Substances Dissolve Answer Key

Howard Stephen Stoker



Section How Substances Dissolve Answer Key:

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water quality and the characteristics of water and water bodies water sciences water pollution and water wastewater treatment Part IV deals with soil science and emphasizes soil as a natural resource highlighting the many interactions between soil and other components of the ecosystem Part V is devoted to showing how decisions regarding handling solid and hazardous waste have or can have profound impact on the environment and the three media discussed in this text air water and soil Finally the epilogue looks at the state of the environment past present and future The emphasis in this brief unit is on mitigating present and future environmental concerns by incorporating technology into the remediation process not by blaming technology for the problem

Scientifica David Sang, Peter Ellis, Derek McMonagle, 2004 Bring your science lessons to life with Scientifica Providing just the right proportion of reading versus doing these engaging resources are differentiated to support and challenge pupils of varying abilities

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Chemistry Trace Jordan, Neville R. Kallenbach, 2017 Chemistry The Molecules of Life offers chemical insights within the context of health pharmaceuticals and the function of biological molecules The contextualized presentation of topics gives students a broad introduction to chemistry and helps them to see the relevance of chemistry to their personal lives

Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

Ebook: Physical Science Tillery, 2016-04-16 Ebook Physical Science

The Freezing Point Eleanor Hawking, AI, 2025-02-12 The Freezing Point explores the physics governing how substances transition from liquid to solid emphasizing that it's more than just a fixed temperature The book highlights the interplay of temperature pressure and

intermolecular forces revealing how manipulating these factors can alter freezing points for various technological applications Understanding freezing is crucial impacting fields from materials science to climate science The book starts by establishing a foundation in thermodynamics covering concepts like enthalpy and entropy It then digs into intermolecular forces explaining how different types of bonding affect freezing temperature A significant portion focuses on pressure s effect on the freezing point often overlooked but essential for understanding deep sea environments and industrial processes Examples of phase diagrams are used to clarify the physics and colligative properties are discussed in the context of antifreeze solutions The book supports its arguments with experimental data phase diagrams and thermodynamic calculations drawing from research articles and industrial case studies Its unique value lies in bridging the gap between theory and practice showcasing real world applications in food preservation cryosurgery and antifreeze design The book progresses systematically building from basic principles to complex applications making it accessible to students and anyone curious about the physics of freezing

Environmental Science Frank R. Spellman, Melissa L. Stoudt, 2013-02-14

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Chemistry John S. Phillips, Cheryl Wistrom, 2000

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