

## 9.1

## THE ARITHMETIC OF EQUATIONS

## SECTION REVIEW

**Objectives**

- Calculate the amount of reactants required or product formed in a chemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

**Key Terms**

- stoichiometry

**Part A Completion**

Copy this completion exercise for class. Your understanding of the concepts and terms is clear and reinforced in this exercise. Write down the completed work at home, school, or home first.

- The coefficients in a balanced chemical equation indicate \_\_\_\_\_ 1. \_\_\_\_\_ the relative number of \_\_\_\_\_ 2. \_\_\_\_\_ of reactants and products. All \_\_\_\_\_ 3. \_\_\_\_\_ stoichiometric calculations begin with a \_\_\_\_\_ 4. \_\_\_\_\_ On a \_\_\_\_\_ 5. \_\_\_\_\_ and \_\_\_\_\_ 6. \_\_\_\_\_ are concerned in every reaction: moles, volumes, and representative particles may not be.
- In solving stoichiometric problems, conversion factors relating moles of reactants to \_\_\_\_\_ 7. \_\_\_\_\_ of products are used.
- If you assume \_\_\_\_\_ 8. \_\_\_\_\_, the equation also tells you about the volume of gas.

**Part B True-False**

Quickly mark off the statements as to be always true, AC: sometimes true, AT: probably true, NT.

- \_\_\_\_\_ 7. The coefficients in a balanced chemical equation can be used to determine the relative limiting reactants to products.
- \_\_\_\_\_ 8. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable units of volume.
- \_\_\_\_\_ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

# Section 9 1 The Arithmetic Of Equations

**Eastern Michigan University, Michigan  
State Normal College**



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