

Chapter 9 Carbon Chemistry

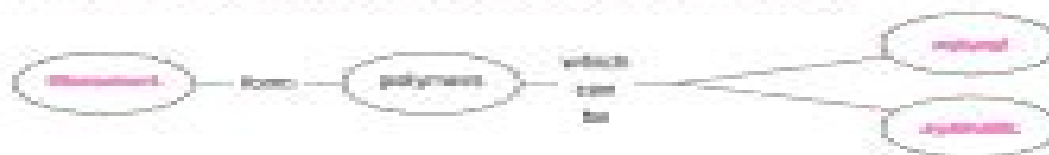
Section 9.3 Polymers

(pages 276–280)

This section explains how polymers form. It also discusses examples of synthetic and natural polymers.

Reading Strategy (page 276)

Identifying Main Ideas As you read, complete the concept map to summarize two main ideas about polymers. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.



- Describe a polymer. A polymer is a large molecule that forms when many smaller molecules are linked together by covalent bonds.
- The smaller molecules that join together to form a polymer are called monomers.
- Is the following sentence true or false? More than one type of monomer can be present in some polymers. True
- Name the two general classifications of polymers.
 - Natural
 - Synthetic

Synthetic Polymers (page 276)

- Name three polymers that can be synthesized.
 - Polyethylene
 - Nylon
 - Rubber
- Is the following sentence true or false? The more carbon atoms there are in a polyethylene chain, the harder the polymer is. True

Natural Polymers (pages 276–280)

- Name four types of polymers that are produced in plant and animal cells.
 - Starches
 - Complex acids
 - Cellulose
 - Proteins
- Circle the letter of the molecular formula of a simple sugar.
 - CH_2O
 - $\text{C}_6\text{H}_{12}\text{O}_6$
 - $\text{C}_{12}\text{H}_{22}\text{O}_{11}$
 - $\text{C}_{10}\text{H}_{16}\text{O}_4$

Section 9 3 Polymers Wordwise

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Table of Contents Section 9 3 Polymers Wordwise

1. Understanding the eBook Section 9 3 Polymers Wordwise
 - The Rise of Digital Reading Section 9 3 Polymers Wordwise
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 9 3 Polymers Wordwise
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 9 3 Polymers Wordwise
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 9 3 Polymers Wordwise
 - Personalized Recommendations
 - Section 9 3 Polymers Wordwise User Reviews and Ratings
 - Section 9 3 Polymers Wordwise and Bestseller Lists
5. Accessing Section 9 3 Polymers Wordwise Free and Paid eBooks
 - Section 9 3 Polymers Wordwise Public Domain eBooks
 - Section 9 3 Polymers Wordwise eBook Subscription Services
 - Section 9 3 Polymers Wordwise Budget-Friendly Options

6. Navigating Section 9 3 Polymers Wordwise eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 9 3 Polymers Wordwise Compatibility with Devices
 - Section 9 3 Polymers Wordwise Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 9 3 Polymers Wordwise
 - Highlighting and Note-Taking Section 9 3 Polymers Wordwise
 - Interactive Elements Section 9 3 Polymers Wordwise
8. Staying Engaged with Section 9 3 Polymers Wordwise
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 9 3 Polymers Wordwise
9. Balancing eBooks and Physical Books Section 9 3 Polymers Wordwise
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 9 3 Polymers Wordwise
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 9 3 Polymers Wordwise
 - Setting Reading Goals Section 9 3 Polymers Wordwise
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 9 3 Polymers Wordwise
 - Fact-Checking eBook Content of Section 9 3 Polymers Wordwise
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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