



## Percentage Yield, *continued*

### Sample Problem H Solution



Given:

mass of  $\text{C}_6\text{H}_6 = 36.8 \text{ g}$

mass of  $\text{Cl}_2 = \text{excess}$

actual yield of  $\text{C}_6\text{H}_5\text{Cl} = 38.8 \text{ g}$

Unknown: percentage yield of  $\text{C}_6\text{H}_5\text{Cl}$

Solution:

Theoretical yield

$$\text{g C}_6\text{H}_6 \times \frac{\text{molar mass factor}}{\text{g C}_6\text{H}_6} \times \frac{\text{mol ratio}}{\text{mol C}_6\text{H}_6} \times \frac{\text{molar mass}}{\text{mol C}_6\text{H}_5\text{Cl}} = \text{g C}_6\text{H}_5\text{Cl}$$

*mol C<sub>6</sub>H<sub>5</sub>Cl*



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Resources



# Section Quiz Limiting Reactants And Percentage Yield

**Victor M. Corman**



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