

Boyle's Law

1. The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
2. Boyle's
3. Boyle's Law

Boyle's Theory and Boyle's Law

1. Boyle's of gases is caused by the elasticity of the gas molecules themselves that collide.
2. If the gas is compressible to small the volume is small which increases the pressure which is small.
3. Force exerts pressure on closed on the walls of the container.
4. Pressure of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a volume of 400 ml. above the pressure is 1000 mm. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 1000 \text{ mm} \times 400 \text{ ml} = P_2 \times 800 \text{ ml} = P_2$$

$$P_2 = \frac{1000 \times 400}{800} = 500 \text{ mm}$$

Ex. 2. A gas is pressure of 100 mm. and 400 ml. of gas has a volume of 800 ml. If the pressure is increased until the volume is 200 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ mm} \times 400 \text{ ml} = P_2 \times 200 \text{ ml} = P_2$$

$$P_2 = \frac{100 \times 400}{200} = 200 \text{ mm}$$

Note Taking Guide Episode 902 Chemistry

Mark Steyvers



Note Taking Guide Episode 902 Chemistry:

Note Taking Guide Episode 902 Chemistry Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Note Taking Guide Episode 902 Chemistry**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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