91171





## Level 2 Physics, 2012

### 91171 Demonstrate understanding of mechanics

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gases after all the oxygen is removed from a sample of air in a glass tube and the temperature remains constant the pressure exerted by the remaining air a does not change b

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web temperature and pressure hints density g l use the ideal gas law to solve for moles first a 0 35 g l b 2 14 g l c 2 86 g l first use the ideal gas law and recall that density is g ll 101 3 x 1 l n x 8 31 x 273 therefore n 0 045 moles density is g l so convert moles to grams by multiplying by the molar mass of so 2

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web answer as temperature of a gas increases pressure will also increase based on the ideal gas law the volume of the tire can only expand so much before the rubber gives and releases the build up of pressure

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web the observed behavior of gases embodied in the empirical gas laws leads to a series of equations that can be summarized by a single equation of state called the ideal gas law equation this shows the relationship between a gas s pressure p temperature t volume v and amount in moles n

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web each gas in a mixture of gases exerts its own pressure independently of the other gases present therefore the pressure of each gas within a mixture is called the partial pressure of the gas dalton s law of partial pressures states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the component gases gas laws quiz chemquiz net - Jun 19 2023

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web at very high pressures a gas will occupy a larger volume than predicted by the ideal gas law which of the following statements is true at low temperatures intermolecular forces become important and the pressure of a gas will be lower than

predicted by