

# Chemistry

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(Chapter 5) (Coordination Compound) (Intext Questions)

(Class 12)

Therefore,  $\mu = \sqrt{n(n+2)} = \sqrt{1(1+2)} = \sqrt{3} = 1.732 \text{ BM}$

On the other hand,  $\text{H}_2\text{O}$  is a weak field ligand. Therefore, it cannot cause the pairing of electrons. This means that the number of unpaired electrons is 5.





Therefore,  $\mu = \sqrt{n(n+2)} = \sqrt{5(5+2)} = \sqrt{35} = 6 \text{ BM}$

Thus, it is evident that  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$  is strongly paramagnetic, while  $[\text{Fe}(\text{CN})_6]^{3-}$  is weakly paramagnetic.

## Question 5.8:

Explain  $[\text{Co}(\text{NH}_3)_6]^{3+}$  is an inner orbital complex whereas  $[\text{Ni}(\text{NH}_3)_6]^{2+}$  is an outer orbital complex.

Answer 5.8:

$[\text{Co}(\text{NH}_3)_6]^{3+}$	$[\text{Ni}(\text{NH}_3)_6]^{2+}$
Oxidation state of cobalt = +3	Oxidation state of Ni = +2
Electronic configuration of cobalt = $d^6$	Electronic configuration of nickel = $d^8$
 <p>NH<sub>3</sub> being a strong field ligand causes the pairing. Therefore, Ni can undergo <math>d^2sp^3</math> hybridization.</p>  <p>Hence, it is an inner orbital complex.</p>	 <p>If NH<sub>3</sub> causes the pairing, then only one 3d orbital is empty. Thus, it can not undergo <math>d^2sp^3</math> hybridization. Therefore, it undergoes <math>sp^3d^2</math> hybridization.</p>  <p>Hence, it forms an outer orbital complex.</p>

## Question 5.9:

Predict the number of unpaired electrons in the square planar  $[\text{Pt}(\text{CN})_4]^{2-}$  ion.

Answer 5.9:

$[\text{Pt}(\text{CN})_4]^{2-}$

In this complex, Pt is in the +2 state. It forms a square planar structure. This means that it undergoes  $dsp^2$  hybridization. Now, the electronic configuration of Pd(+2) is  $5d^8$ .



CN<sup>-</sup> being a strong field ligand causes the pairing of unpaired electrons. Hence, there are no unpaired electrons in  $[\text{Pt}(\text{CN})_4]^{2-}$ .

## Question 5.10:

The hexaquo manganese (II) ion contains five unpaired electrons, while the hexacyanoion contains only one unpaired electron. Explain using Crystal Field Theory.

Answer 5.10:

$[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$	$[\text{Mn}(\text{CN})_6]^{4-}$
Mn is in the +2 oxidation state.	Mn is in the +2 oxidation state.
The electronic configuration is $d^5$	The electronic configuration is $d^5$ .
The crystal field is octahedral. Water is a weak field ligand. Therefore, the arrangement of the electrons in $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ is $t_{2g}^3e_g^2$ .	The crystal field is octahedral. Cyanide is a strong field ligand. Therefore, the arrangement of the electrons in $[\text{Mn}(\text{CN})_6]^{4-}$ is $t_{2g}^5e_g^0$ .
Hence, hexaquo manganese (II) ion has five unpaired electrons	While hexacyano ion has only one unpaired electron.

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