

NCERT Solutions for Class 11

Chemistry

Chapter – 1 Some basic concepts of Chemistry

1. Calculate the molecular mass of the following:

i. H_2O

Ans: The molecular mass of water, H_2O , can be calculated by the following steps given below:

= (2 $\times$ Atomic mass of hydrogen element) + (1 $\times$ Atomic mass of oxygen element)

= [(2 $\times$ 1.0084 u) + (1 $\times$ 16.00 u)]

= 2.016 u + 16.00 u

= 18.016 u or 18.02 u

ii. CO_2

Ans: The molecular mass of carbon dioxide, CO_2 , is calculated below:

= (1 $\times$ Atomic mass of carbon) + (2 $\times$ Atomic mass of oxygen)

= [(1 $\times$ 12.011 u) + (2 $\times$ 16.00 u)]

= 12.011 u + 32.00 u

= 44.01 u

iii. CH_4

Ans: The molecular mass of methane, CH_4 , is calculated below in step by step manner:

= (1 $\times$ Atomic mass of carbon) + (4 $\times$ Atomic mass of hydrogen)

= [(1 $\times$ 12.011 u) + (4 $\times$ 1.008 u)]

= 12.011 u + 4.032 u

= 16.043 u

2. Calculate the mass percent of different elements present in sodium sulphate (Na_2SO_4).

Ans: The given compound in the question is sodium sulphate and its formula is Na_2SO_4 . Its molecular formula is calculated below:

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Y Pai

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