

# Chemistry

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(Chapter 1)(Solutions) (Exercises)  
(Class 12)

## Question 1.1:

Define the term solution. How many types of solutions are formed? Write briefly about each type with an example.

## Answer 1.1:

A homogeneous mixture of two or more non-reacting substances whose composition varies within certain fixed limits, known as a solution. Generally, the component that is present in the largest quantity is known as *solvent*. Solvent determines the physical state in which solution exists. One or more components present in the solution other than solvent are called solutes.

When the solutions contains two, three or four components, it is known as binary, ternary or quaternary solutions respectively. On the basis of physical state of components, the solutions are of the following types:

- 1) When both solute and solvent are in solid state. Example: Alloys.
- 2) When solute is in solid and solvent is in liquid state. Example: Sugar or salt solutions.
- 3) When solute is in solid and solvent is in gaseous state. Example: iodine vapours in air.
- 4) When both solute and solvent are in liquid state. Example: alcohol in water.
- 5) When solute is in liquid and solvent is in solid state. Example: Zinc amalgam.
- 6) When solute is in liquid and solvent is in gaseous state. Example: water vapour in air.
- 7) When both solute and solvent are in gaseous state. Example: air.
- 8) When solute is in gaseous and solvent is in liquid state. Example: aerated drinks.
- 9) When solute is in gaseous and solvent is in solid state. Example: dissolve gas in mineral.

## Question 1.2:

Give an example of a solid solution in which the solute is a gas.

## Answer 1.2:

Dissolved gases in minerals.

## Question 1.3:

Define the following terms:

- (i) Mole fraction                      (ii) Molality                      (iii) Molarity                      (iv) Mass percentage.

## Answer 1.3:

(i) Mole fraction: Commonly used symbol for mole fraction is  $x$  and subscript used on the right hand side of  $x$  denotes the component. It is defined as:

$$\text{Mole fraction of a component} = \frac{\text{Number of moles of the component}}{\text{Number of moles of all the component}}$$

(ii) Molality: Molality ( $m$ ) is defined as the number of moles of the solute per kilogram (kg) of the solvent and is expressed as:

$$\text{Molality (m)} = \frac{\text{Moles of solute}}{\text{Mass of solvent in kg}}$$

(iii) Molarity: Molarity ( $M$ ) is defined as number of moles of solute dissolved in one litre (or one cubic decimetre) of solution.

$$\text{Molarity (M)} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$$

(iv) Mass percentage: The mass percentage of a component of a solution is defined as:

$$\text{Mass \% of a component} = \frac{\text{Mass of the component in the solution}}{\text{Total mass of the solution}} \times 100$$

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