

Thin Layer Chromatography

I. Purpose

Thin Layer Chromatography (TLC) is a laboratory technique used to identify different compounds present within a mixture. There are two phases: The stationary phase, which contains a solid and usually polar, and the mobile phase, which can be either a liquid or gas which can be a pure solvent or a mixture of solvents.¹ The TLC plates consist of a solid surface such as glass, metal or plastic with a thin adsorbent later on the surface and provides the stationary phase. In the lab, silica gel will be act as the solid, stationary phase and 99:1 ethyl acetate. The liquid, mobile phase will be represented by the use of acetic acid. If the compound in question forms dark spots on the TLC plate, then it will stick to the stationary phase because it is polar. They tend to form spots close to the starting, origin line of the TLC plate. On the other hand, the spots will flow with the mobile phase because the compound is less polar compared to the other compound. Nonpolar substances will also form spots that traveled a further distance compared to those that are slightly polar. A more polar compound will tend to strongly bind to the silica gel compared a compound less polar. To compare the distances between each substance, a retention factor (R_f) is calculated for each spot. It is determined by the ratio traveled by the compound divided to the distance the solvent has traveled.¹ An R_f value can be enlarged by raising the polarity of the solvent, while it can also be decreased by increasing the polarity of the compound.²

II. Safety

Safety Data Sheet – NFPA ^{3,4}				
	Acetaminophen ⁵	Aspirin/Acetylsalicylic Acid ⁶	Dichloromethane ⁷	Ethyl Alcohol ⁸
Health	2	1	2	3
Flammability	1	1	1	2
Corrosivity	0	1	0	0
Other Hazards	N/A	0	N/A	0

Table 1: Safety Data Sheet

III. Experiment⁴

1. Obtain a TLC Plate of the appropriate size that can fit into the developing chamber.

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