

## **Full Papers**

#### Modeling of Phosphoric Acid Purification by Liquid-Liquid Extraction

By Bendada Ahlem, Meniai Ahdeslam-Hassen\*, and Bencheikh Lehocine Mossaab

Phosphoric acid is an important industrial chemical. It is mainly produced by the wet process which consists of an attack of the phosphate rock by sulphuric acid leading to a complex solution containing a large number of impurities such as metal ions like Fe<sup>5+</sup>, Al<sup>5+</sup>. Mg<sup>5+</sup>, Cd<sup>5+</sup> etc., which can effectively be recovered by solvent extraction process. In this present work a model of the liquid-liquid extraction process is presented. It is mainly based on thermodynamics, where two different routes have been tested for the modeling of the complexation process. The method has been tested using a model system with Al<sup>5+</sup> as the contaminant, dodocyl naphthalene sulphonic acid (HDDNSA) as the complexing agent and kerosene as the diluent. The study has also investigated the influence of various parameters such as the pH of the aqueous phase and the initial concentration of the phosphoric acid.

#### 1 Introduction

Phosphoric acid is an important industrial chemical, which has several important uses such as an intermediate for the fertiliser industry, for surface treatment in the metal industry and also as an additive in the food industry. It is mainly produced by two different routes: the wet processes or the thermal processes, with the former being the most widely used. It is known as the WPA process and it consists of an attack by sulphuric acid of the phosphate rock, which is the principal source of the dissolved and suspended impurities. This will lead to a complex solution containing a large number of organic and inorganic species, particularly metal ions, the chief ones being Fe3-, Al3-, Mg2-, Cd2- etc. Consequently, the presence of these impurities can induce serious limitations for the use of the phosphoric acid and hence one can see the importance of efficient purification processes such as solvent extraction, which can effectively recover the ionic species.

In this work, an attempt to model the phosphoric acid purification process by liquid-liquid extraction is presented.

#### 2 Description of the Model

The proposed approach is based on a complex formation between ionic species present in the solution, using a judiciously selected complexing agent diluted in an inert organic solvent such as kerosene, chloroform etc., followed by a phase separation, where a part of the amount of the complexes formed migrate into the organic liquid layer, to ensure phase equilibria.

$$H_0PO_4 \leftrightarrow H_0PO_4^- + H^+$$
  $K_0 = 7.107 \ 10^{-3}$  (R1)

$$H_2PO_4^- \leftrightarrow HPO_4^{2-} + H^+ \qquad K_2 = 4.275 \ 10^{-2}$$
 (R2)

$$HPO_4^{3-} \rightarrow PO_4^{3-} + H^{\circ}$$
  $K_3 = 1.15 \cdot 10^{-12}$  (R3)

$$H_3PO_4 + H_2PO_4^- \leftrightarrow H_3P_2O_3^ K_4 = 1.263$$
 (R4)

$$H' + H_5P_2O_8 \rightarrow H_6P_2O_8$$
  $K_5 = 0.3$  (R5)

$$H_2O \leftrightarrow H^+ + OH^ K_W = 10^{-14}$$
 (R6)

The K's express the dissociation or hydrolysis reactions equilibrium constants. In this work, the additional reaction is that between the metallic cation and the selected complexing agent which is the HDDNSA for the contamination with AI (III). For this case only one single reaction has been determined from experimental measurements, as reported in [1], which is expressed as follows:

$$AI^{1+} + \frac{3}{3.5}(HR)_{3.5} \rightarrow AIR_3 + 3H^+ K_6 = 52.7$$
 (R7)

R denotes the DDNSA radical.

During this process and at T = 298 K, the following coupled equilibria take place<sup>13</sup>:

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<sup>)</sup> List of symbols at the end of the paper.

# **Papers Of Production Phosphoric Acid**

**Daniel F McAuley** 

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#### **Table of Contents Papers Of Production Phosphoric Acid**

- 1. Understanding the eBook Papers Of Production Phosphoric Acid
  - The Rise of Digital Reading Papers Of Production Phosphoric Acid
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Papers Of Production Phosphoric Acid
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - o Features to Look for in an Papers Of Production Phosphoric Acid
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Papers Of Production Phosphoric Acid
  - Personalized Recommendations
  - Papers Of Production Phosphoric Acid User Reviews and Ratings
  - Papers Of Production Phosphoric Acid and Bestseller Lists
- 5. Accessing Papers Of Production Phosphoric Acid Free and Paid eBooks
  - Papers Of Production Phosphoric Acid Public Domain eBooks
  - Papers Of Production Phosphoric Acid eBook Subscription Services
  - Papers Of Production Phosphoric Acid Budget-Friendly Options
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- o ePub, PDF, MOBI, and More
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- Papers Of Production Phosphoric Acid Enhanced eBook Features
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- 8. Staying Engaged with Papers Of Production Phosphoric Acid
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Papers Of Production Phosphoric Acid
- 9. Balancing eBooks and Physical Books Papers Of Production Phosphoric Acid
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Papers Of Production Phosphoric Acid
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Papers Of Production Phosphoric Acid
  - Setting Reading Goals Papers Of Production Phosphoric Acid
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Papers Of Production Phosphoric Acid
  - Fact-Checking eBook Content of Papers Of Production Phosphoric Acid
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
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  - Interactive and Gamified eBooks

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