

Date: _____

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Fission and Fusion

11. The number and mass of nuclei changes in a fission process. The number _____, whereas the mass _____.
12. How is the energy released? How about is the difference between the mass of the separate particles before and the products in a fission reaction?
13. Calculate the energy released in the fission of two nuclei and write down the mass of the fission products. How is the energy released? How about is the difference between the mass of the separate particles before and the products in a fission reaction?
14. Calculate the energy released in the fission of two nuclei and write down the mass of the fission products. How is the energy released? How about is the difference between the mass of the separate particles before and the products in a fission reaction?

$$E = mc^2 = (2.014102 \text{ u} + 1.008665 \text{ u} - 2.013546 \text{ u} - 1.008665 \text{ u}) \times 931.5 \text{ MeV/u} = 200.9 \text{ MeV}$$

$$E = mc^2 = (2.014102 \text{ u} + 1.008665 \text{ u} - 2.013546 \text{ u} - 1.008665 \text{ u}) \times 931.5 \text{ MeV/u} = 200.9 \text{ MeV}$$

15. Calculate the energy released in the following fission reactions. Calculate the energy released in the fission of two nuclei and write down the mass of the fission products.

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Barium-141} + \text{Krypton-92} + 3\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Strontium-90} + \text{Zirconium-144} + 2\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Iodine-131} + \text{Xenon-142} + 3\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Cesium-137} + \text{Rubidium-94} + 2\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Iodine-131} + \text{Xenon-142} + 3\text{neutrons} + 200 \text{ MeV} \quad (\text{showing energy released})$$

16. Calculate the energy released in the following fission reactions. Calculate the energy released in the fission of two nuclei and write down the mass of the fission products.

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Barium-141} + \text{Krypton-92} + 3\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Strontium-90} + \text{Zirconium-144} + 2\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Iodine-131} + \text{Xenon-142} + 3\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Cesium-137} + \text{Rubidium-94} + 2\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Iodine-131} + \text{Xenon-142} + 3\text{neutrons}$$

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Cesium-137} + \text{Rubidium-94} + 2\text{neutrons}$$

17. Calculate the energy released in the following fission reactions. Calculate the energy released in the fission of two nuclei and write down the mass of the fission products.

$$\text{Uranium-235} + \text{neutron} \rightarrow \text{Barium-141} + \text{Krypton-92} + 3\text{neutrons} + 200 \text{ MeV} \quad (\text{showing energy released})$$

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Michael Seilmaier



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