



2

ATOMIC THEORY: ISOTOPES

Solutions

Complete the following table.

Isotope Name		Atomic Number	Mass Number	Symbol	Number of Protons	Number of Neutrons
1.	carbon-14	6	14	$^{14}_6\text{C}$	6	8
2.	oxygen-16	8	16	$^{16}_8\text{O}$	8	8
3.	polonium-212	84	212	$^{212}_{84}\text{Po}$	84	128
4.	uranium-238	92	238	$^{238}_{92}\text{U}$	92	146
5.	hydrogen-2	1	2	^2_1H	1	1
6.	helium-4	2	4	^4_2He	2	2
7.	thorium-232	90	232	$^{232}_{90}\text{Th}$	90	142
8.	carbon-12	6	12	$^{12}_6\text{C}$	6	6
9.	lawrencium-257	103	257	$^{257}_{103}\text{Lr}$	103	154
10.	hydrogen-1	1	1	^1_1H	1	0

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Table of Contents Review Sheet Calculations Test 2 Answer Key

1. Understanding the eBook Review Sheet Calculations Test 2 Answer Key
 - The Rise of Digital Reading Review Sheet Calculations Test 2 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Sheet Calculations Test 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Sheet Calculations Test 2 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Sheet Calculations Test 2 Answer Key
 - Personalized Recommendations
 - Review Sheet Calculations Test 2 Answer Key User Reviews and Ratings
 - Review Sheet Calculations Test 2 Answer Key and Bestseller Lists

5. Accessing Review Sheet Calculations Test 2 Answer Key Free and Paid eBooks
 - Review Sheet Calculations Test 2 Answer Key Public Domain eBooks
 - Review Sheet Calculations Test 2 Answer Key eBook Subscription Services
 - Review Sheet Calculations Test 2 Answer Key Budget-Friendly Options
6. Navigating Review Sheet Calculations Test 2 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Review Sheet Calculations Test 2 Answer Key Compatibility with Devices
 - Review Sheet Calculations Test 2 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Sheet Calculations Test 2 Answer Key
 - Highlighting and Note-Taking Review Sheet Calculations Test 2 Answer Key
 - Interactive Elements Review Sheet Calculations Test 2 Answer Key
8. Staying Engaged with Review Sheet Calculations Test 2 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Sheet Calculations Test 2 Answer Key
9. Balancing eBooks and Physical Books Review Sheet Calculations Test 2 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Sheet Calculations Test 2 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review Sheet Calculations Test 2 Answer Key
 - Setting Reading Goals Review Sheet Calculations Test 2 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Sheet Calculations Test 2 Answer Key
 - Fact-Checking eBook Content of Review Sheet Calculations Test 2 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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