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## Unit 4: Gases Unit Review Sheet

Chemistry R

1. What are the 4 major ideas of Kinetic Molecular Theory?

- A) Random, straight line motion, constant
- B) elastic collisions
- C) volume of particles is insignificant
- D) no attractive forces, no IMF

2. Ideal vs. Real Gases

PLIGHT He, H<sub>2</sub>

- no attractive forces
- light & small gases

3. Gas Laws

A) Boyle's:  $P \downarrow, V \uparrow$  ( $P \uparrow, V \downarrow$ )

Formula:  $P_1 V_1 = P_2 V_2$

- Marshmallow/Peep Experiment

( $\downarrow P$  around the peep, so  $P$  inside was now greater than  $P$  outside of peep, so  $V$  of peep  $\uparrow$ , peep will expand)

B) Charles:  $V \uparrow, T \uparrow$  ( $T \downarrow, V \downarrow$ )

Formula: \_\_\_\_\_

- > Balloon in the freezer example

$\downarrow$   $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

( $\downarrow T$ , the molecules slow down which means fewer collisions with the sides of the balloon, so not much pressure inside so  $V \downarrow$ )

# Review Gases Section 1 Answer Key Holt

**Puja Mehta**



## **Review Gases Section 1 Answer Key Holt:**

## Reviewing **Review Gases Section 1 Answer Key Holt**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Review Gases Section 1 Answer Key Holt**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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