

## CHAPTER 9 REVIEW

*Stoichiometry*

## MIXED REVIEW

**SHORT ANSWER** Answer the following questions in the space provided.

1. Given the following equation:
- $\text{C}_3\text{H}_8(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

\_\_\_\_\_ <sup>4</sup> a. What is the value of the coefficient  $x$  in this equation?\_\_\_\_\_ <sup>40.0 g/mol</sup> b. What is the molar mass of  $\text{C}_3\text{H}_8$ ?\_\_\_\_\_ <sup>4 mol O<sub>2</sub></sup>  
\_\_\_\_\_ <sup>2 mol H<sub>2</sub>O</sup> c. What is the mole ratio of  $\text{O}_2$  to  $\text{H}_2\text{O}$  in the above equation?\_\_\_\_\_ <sup>0.20 mol</sup> d. How many moles are in an 8.0 g sample of  $\text{C}_3\text{H}_8$ ?\_\_\_\_\_ <sup>32</sup> e. If  $x$  mol of  $\text{C}_3\text{H}_8$  react, how many moles of  $\text{CO}_2$  are produced, in terms of  $x$ ?  $\text{mol CO}_2 = (x \text{ mol}) \left( \frac{3 \text{ mol CO}_2}{1 \text{ C}_3\text{H}_8} \right) = 3x$ 

2. a. What is meant by
- ideal conditions*
- relative to stoichiometric calculations?

\_\_\_\_\_ <sup>STP  $\Rightarrow$  101.3 kPa</sup>  
\_\_\_\_\_ <sup>0°C</sup>

- b. What function do ideal stoichiometric calculations serve?

\_\_\_\_\_ <sup>allow us to use known/constant values</sup>

- c. Are actual yields typically larger or smaller than theoretical yields?

\_\_\_\_\_ <sup>smaller</sup>**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

3. Assume the reaction represented by the following equation goes all the way to completion:



- \_\_\_\_\_
- <sup>4 mol H<sub>2</sub></sup>
- a. If 6 mol of
- $\text{H}_2$
- are consumed, how many moles of
- $\text{NH}_3$
- are produced?

$$\text{mol NH}_3 = (6 \text{ mol H}_2) \left( \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 4 \text{ mol}$$

- \_\_\_\_\_
- <sup>8.5 g NH<sub>3</sub></sup>
- b. How many grams are in a sample of
- $\text{NH}_3$
- that contains
- $3.0 \times 10^{23}$
- molecules?

$$g \text{ NH}_3 = (3.0 \times 10^{23} \text{ molecules}) \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) \left( \frac{17.0 \text{ g}}{\text{mol}} \right)$$

# Review Modern Chemistry Answer Sheet

**David W. Oxtoby, H. Pat Gillis, Laurie J.  
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## **Table of Contents Review Modern Chemistry Answer Sheet**

1. Understanding the eBook Review Modern Chemistry Answer Sheet
  - The Rise of Digital Reading Review Modern Chemistry Answer Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Review Modern Chemistry Answer Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Review Modern Chemistry Answer Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Review Modern Chemistry Answer Sheet
  - Personalized Recommendations
  - Review Modern Chemistry Answer Sheet User Reviews and Ratings
  - Review Modern Chemistry Answer Sheet and Bestseller Lists
5. Accessing Review Modern Chemistry Answer Sheet Free and Paid eBooks

- Review Modern Chemistry Answer Sheet Public Domain eBooks
- Review Modern Chemistry Answer Sheet eBook Subscription Services
- Review Modern Chemistry Answer Sheet Budget-Friendly Options
- 6. Navigating Review Modern Chemistry Answer Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Review Modern Chemistry Answer Sheet Compatibility with Devices
  - Review Modern Chemistry Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Review Modern Chemistry Answer Sheet
  - Highlighting and Note-Taking Review Modern Chemistry Answer Sheet
  - Interactive Elements Review Modern Chemistry Answer Sheet
- 8. Staying Engaged with Review Modern Chemistry Answer Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Review Modern Chemistry Answer Sheet
- 9. Balancing eBooks and Physical Books Review Modern Chemistry Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Review Modern Chemistry Answer Sheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Review Modern Chemistry Answer Sheet
  - Setting Reading Goals Review Modern Chemistry Answer Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Modern Chemistry Answer Sheet
  - Fact-Checking eBook Content of Review Modern Chemistry Answer Sheet
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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