

**MARK SCHEME for the May/June 2014 series**

**0620/11**

**0620 CHEMISTRY**

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# Paper Igcse Chemistry May June 2014 0620 11

**Daniel F McAuley**



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web a rate law shows how the rate of a chemical reaction depends on reactant concentration for a reaction such as  $a \rightarrow b$  products the rate law generally has the form  $\text{rate} = k [a]^n$  where  $k$  is a proportionality constant called the rate constant and  $n$  is the order of the reaction with respect to  $a$  the value of  $n$  is not related to the reaction stoichiometry and



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web the rate determining slower step gives a rate law showing second order dependence on the  $\text{NO}_2$  concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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web rate laws and reaction order the relation between the rate of a reaction and the concentrations of reactants is expressed by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$  has been found to be directly proportional to the concentration of  $\text{N}_2\text{O}_5$  text rate  $k$   $\text{N}_2\text{O}_5$

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text  $\text{O}_2$  text  $\text{C}$  see table below notice that the starting concentrations of  $\text{C}_2\text{H}_4$  and  $\text{C}_2\text{H}_6$

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hydrogen peroxide  $\text{H}_2\text{O}_2$  were used in order to calculate the rate of reaction  
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web rate laws sometimes called differential rate laws or rate equations are mathematical expressions that describe the relationship between the rate of a chemical reaction and the concentration of its reactants as an example consider the reaction described by the chemical equation  $\text{a} + \text{b} \rightarrow \text{products}$

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web the determination of a rate law is a matter of finding an empirical equation that adequately describes reaction rate data we can distinguish two general approaches to this task one approach is to measure reaction rate directly that is for  $\text{a} + \text{b} \rightarrow \text{c}$  we measure the reaction rate in experiments where the concentrations  $\text{a}$  and  $\text{b}$

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to be text rate  $k$  left  $\text{CH}_3\text{OH}$  right nonumber what is the order of reaction with respect to methanol and ethyl acetate and what is the overall order of reaction

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web sep 22 2021 the equilibrium system can be written in the general form  $ce\ ha\ aq\ h\ aq\ a\ aq$  label 7 the equilibrium constant expression for this reaction is  $k = \frac{ce\ h\ ce\ a\ ce\ ha}{a}$  label 8 where we denote the equilibrium constant  $k$  with a subscript  $a$  for acid

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web equilibrium moles  $hscn\ 2.00 \times 10^{-5}\ mol\ 3.00 \times 10^{-6}\ mol\ 1.70 \times 10^{-5}\ mol\ hscn$  knowing the number of moles of  $fe^{3+}$  and  $hscn$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $mol\ fe^{3+}\ 3.170 \times 10^{-5}\ mol\ fe^{3+}\ 8.50 \times 10^{-4}\ m$

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